

Name: _____ Date: _____ Period: _____

What happened to
Avogadro when he
got bit by 6.02×10^{23}
mosquitoes?
He got Mol-aria

POGIL- Stoichiometry

How do chemists use balanced chemical equations?

Why?

Chemists use balanced chemical equations as a basis to calculate how much reactant is needed or product is formed in a reaction. This is called **Stoichiometry- (stoi-key-ah-meh-tree)** Another way of looking at it is using the mole ratio from the balanced equation and information about one compound in the reaction to determine information about another compound in the equation. A mole ratio is a conversion factor derived from the coefficients of a balanced chemical equation interpreted in terms of moles. In chemical calculations, mole ratios are used to convert between moles one thing and moles of another.

Purpose: In this activity we will address the question: *How do I convert between different chemical species in a given reaction?*

MODEL 1: Mole-to-Mole Stoichiometry (1-step)

Example: mole-to-mole

Question: Given the following equation:



How many moles of H_2 are needed to produce 34.8 moles of NH_3 ?

Answer:

$$\frac{34.8 \text{ mol NH}_3}{1} \times \frac{3 \text{ mol H}_2}{2 \text{ mol NH}_3} = 52.2 \text{ mol H}_2$$

OR

$$34.8 \text{ mol NH}_3 \times \frac{3 \text{ mol H}_2}{2 \text{ mol NH}_3} = 52.2 \text{ mol H}_2$$

Key Questions:

1. What was the given in the example above?
2. Where did the $\frac{3 \text{ mol H}_2}{2 \text{ mol NH}_3}$ come from?
3. Double check to see if 52.2 moles of H_2 is the correct answer. SHOW the math below.

Basic Stoichiometry Pogil Activity Solved

David Callaghan



Basic Stoichiometry Pogil Activity Solved:

Chemistry Education in the ICT Age Minu Gupta Bhowon, Sabina Jhaumeer-Laulloo, Henri Li Kam Wah, Ponnadurai Ramasami, 2009-07-21
The 20 International Conference on Chemical Education 20 ICCE which had rd th Chemistry in the ICT Age as the theme was held from 3 to 8 August 2008 at Le Meridien Hotel Pointe aux Piments in Mauritius With more than 200 participants from 40 countries the conference featured 140 oral and 50 poster presentations th Participants of the 20 ICCE were invited to submit full papers and the latter were subjected to peer review The selected accepted papers are collected in this book of proceedings This book of proceedings encloses 39 presentations covering topics ranging from fundamental to applied chemistry such as Arts and Chemistry Education Biochemistry and Biotechnology Chemical Education for Development Chemistry at Secondary Level Chemistry at Tertiary Level Chemistry Teacher Education Chemistry and Society Chemistry Olympiad Context Oriented Chemistry ICT and Chemistry Education Green Chemistry Micro Scale Chemistry Modern Technologies in Chemistry Education Network for Chemistry and Chemical Engineering Education Public Understanding of Chemistry Research in Chemistry Education and Science Education at Elementary Level We would like to thank those who submitted the full papers and the reviewers for their timely help in assessing the papers for publication th We would also like to pay a special tribute to all the sponsors of the 20 ICCE and in particular the Tertiary Education Commission <http://tec.intnet.mu> and the Organisation for the Prohibition of Chemical Weapons <http://www.opcw.org> for kindly agreeing to fund the publication of these proceedings

Research Handbook on Classroom Observation Sean Kelly, 2025-03-12 This comprehensive Research Handbook offers cutting edge insights into classroom observation systems for teacher evaluation and professional development with an emphasis on both applied and discovery oriented inquiry

A Stoichiometry Unit David Callaghan, 2004

Calculating in Chemistry Cosmas C Mbuli, 2022-09-20 This book is intended to help students fully grasp calculations involving reacting quantities Students in higher courses may find it a helpful revision and enhance their clarity The book completely discusses the key topics in basic stoichiometry including the mole concept reacting quantities and empirical and molecular formulas It begins with the basic concepts and formulas required to convert various quantities to moles or amount of substance This particularly useful to a beginner Chapter 2 describes how to calculate reacting reacting quantities and therefore provides a general step by step framework or approach by which to solve these problems The chapter also describes and applies the concept of a limiting reagent Two methods of determining a limiting reagent are explained and illustrated The concept of limiting reagent is extended to reacting volumes of gases with a short cut method A short cut method for solving reacting quantities involving masses and volumes of gases is also given Chapter 3 describes the calculations involved in the practical determination of molecular and empirical formulas A clear meaning of percentage composition of mass is provided and used to solve problems in a step by step manner In chapter 4 we discuss percentage yield and purity A number of examples are given to illustrate how formulas of yield and purity are used in

various circumstances A student will find these examples helpful in relating different formulas for percentage purity The last chapter introduces a graphical method for reacting quantities The method may provide a new way of looking at chemical reactions Examples are given to illustrate the method including how it can be used to determine limiting reagents It is hoped that the book will provide all the necessary knowledge and skills to students studying an introductory chemistry course Teachers may also find this book a good resource for their lessons in stoichiometry

Supporting Traditional Methods with a Constructivist Approach to Learning Kenneth Munoz Abayan, ProQuest Dissertations and Theses (Electronic resource collection), 2014 Stoichiometry is a fundamental topic in chemistry that measures a quantifiable relationship between atoms molecules etc Stoichiometry is usually taught using expository teaching methods Students are passively given information in the hopes they will retain the transmission of information to be able to solve stoichiometry problems masterfully Cognitive science research has shown that this kind of instructional teaching method is not very effecting in meaningful learning practice Instead students must take ownership of their learning The students need to actively construct their own knowledge by receiving interpreting integrating and reorganizing that information into their own mental schemas In the absence of active learning practices tools must be created in such a way to be able to scaffold difficult problems by encoding opportunities necessary to make the construction of knowledge memorable thereby creating a usable knowledge base Using an online e learning tool and its potential to create a dynamic and interactive learning environment may facilitate the learning of stoichiometry The study entailed requests from volunteer students IRB consent form a baseline questionnaire random assignment of treatment pre and post test assessment and post assessment survey These activities were given online A stoichiometry based assessment was given in a proctored examination at the University of Texas at Arlington UTA campus The volunteer students who took part in these studies were at least 18 of age and were enrolled in General Chemistry 1441 at the University of Texas at Arlington Each participant gave their informed consent to use their data in the following study Students were randomly assigned to one of 4 treatments groups based on teaching methodology Dimensional Analysis Operational Method Ratios and Proportions and a control group who just received instruction through lecture only In this study an e learning tool was created to demonstrate several methodologies on how to solve stoichiometry which are all supported by chemical education research Comparisons of student performance based on pre and post test assessment and a stoichiometry based examination was done to determine if the information provided within the e learning tool yielded greater learning outcomes compared to the students in the absence of scaffold learning material The e learning tool was created to help scaffold the problem solving process necessary to help students N 394 solve stoichiometry problems Therein the study investigated possible predictors for success on a stoichiometry based examination students conceptual understanding of solving stoichiometry problems and their explanation of reasoning It was found that the way the student answered a given stoichiometry question i e whether the student used dimensional analysis operational method or any other process was not

statistically relevant $p < 0.05$ More importantly if the students were able to describe their thought process clearly these students scored significantly higher on stoichiometry test mean 84 p

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