

# 7B EXAM PRACTICE

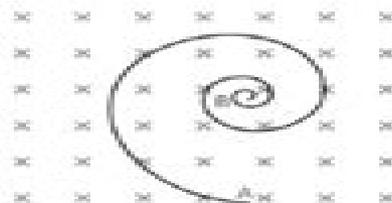
- 1 An electron moving at  $2.05 \times 10^6 \text{ km s}^{-1}$  is made to move in a circular path by a  $0.98 \text{ T}$  magnetic field applied perpendicular to its velocity. The radius of the circle that the electron travels in is:

- A  $1.19 \times 10^{-7} \text{ m}$
- B  $0.119 \text{ mm}$
- C  $8400 \text{ m}$
- D  $1.31 \times 10^{23} \text{ m}$

[1]

(Total for Question 1 = 1 mark)

- 2 The diagram shows the track of a charged particle travelling in a bubble chamber detector in which the magnetic field goes into the page.



The particle:

- A is negative and moves from A to B
- B is positive and moves from A to B
- C is negative and moves from B to A
- D is positive and moves from B to A

[1]

(Total for Question 2 = 1 mark)

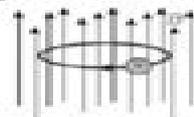
- 3 In order to investigate the structure of the nucleus, high energies are needed. One reason for this is that:

- A The radius of circular motion of a particle in a cyclotron is proportional to the kinetic energy
- B The nucleus can only be ionised by high energy particles.
- C The nucleus will absorb low energy particles.
- D The de Broglie wavelength is inversely proportional to a particle's velocity.

[1]

(Total for Question 3 = 1 mark)

- 4 A strong magnetic field of flux density  $B$  can be used to trap a positive ion by making it follow a circular orbit, as shown.



- (a) Explain how the magnetic field maintains the ion in a circular orbit. You may add to a copy of the diagram above if you wish. [2]

- (b) Show that the mass  $m$  of the ion will be given by

$$m = \frac{Bq}{2\pi f}$$

where  $q$  is the charge on the ion and  $f$  is the number of revolutions per second.

[3]

- (c) The above arrangement will not prevent a positive ion from moving vertically. To do this, a weak electric field is applied using the arrangement shown below.



- (i) Explain how the electric field prevents the ion moving vertically. [2]

- (ii) This device is known as a Penning Trap. It can be used to determine the mass of an ion to an accuracy of 3 parts in 10 million.

Show that the mass of a sulfur ion can be measured to the nearest  $0.00001 \text{ u}$ .

mass of sulfur ion =  $32.0645 \text{ u}$

'u' is an abbreviation for the atomic mass unit.  $1 \text{ u}$  is equivalent to  $931.5 \text{ MeV}$

[2]

- (iii) Under certain conditions, nuclei of sulfur emit a gamma ray with a known energy of  $2.2 \text{ MeV}$ .

Calculate the resulting loss in mass of a sulfur ion in  $\text{u}$  and show that this value could be determined by the Penning Trap technique.

[2]

(Total for Question 4 = 11 marks)

# 7b Practice Form G

**Ian Pickup**



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