

**AP[®] CHEMISTRY
2019 SCORING GUIDELINES**

Question 1



Ethanol, $\text{C}_2\text{H}_5\text{OH}$, will combust in air according to the equation above.

(a) Is $\text{O}_2(g)$ oxidized in the reaction, or is it reduced? Justify your answer in terms of oxidation numbers.

O_2 is reduced; its oxidation number changes from 0 in O_2 to -2 in CO_2 and H_2O .

1 point is earned for the correct answer with an appropriate justification.

(b) When a sample of $\text{C}_2\text{H}_5\text{OH}$ was combusted, the volume of $\text{CO}_2(g)$ produced was 18.0 L when measured at 21.7°C and 1.03 atm. Determine the number of moles of $\text{CO}_2(g)$ that was produced.

$$n = \frac{PV}{RT} = \frac{(1.03 \text{ atm})(18.0 \text{ L})}{(0.08206 \text{ L} \cdot \text{atm mol}^{-1} \text{ K}^{-1})(273.15 + 21.7) \text{ (K)}} = 0.766 \text{ mol CO}_2$$

1 point is earned for the correct calculation.

(c) Determine the volume of $\text{C}_2\text{H}_5\text{OH}(l)$, in mL, that was combusted to produce the volume of $\text{CO}_2(g)$ collected in part (b). (The density of $\text{C}_2\text{H}_5\text{OH}(l)$ is 0.79 g/mL .)

$$0.766 \text{ mol CO}_2 \times \frac{1 \text{ mol C}_2\text{H}_5\text{OH}}{2 \text{ mol CO}_2} \times \frac{46.07 \text{ g C}_2\text{H}_5\text{OH}}{1 \text{ mol C}_2\text{H}_5\text{OH}} \times \frac{1 \text{ mL}}{0.79 \text{ g}} = 22 \text{ mL}$$

1 point is earned for the correct stoichiometry for moles of $\text{C}_2\text{H}_5\text{OH}$ and CO_2 .

1 point is earned for the correct volume of $\text{C}_2\text{H}_5\text{OH}$.

(d) Determine the amount of heat, in kJ, that is released by the combustion reaction.

$$q = n \Delta H_{\text{comb}}^\circ$$

$$q = 0.766 \text{ mol CO}_2 \times \frac{1 \text{ mol}_{\text{rxn}}}{2 \text{ mol CO}_2} \times \frac{-1270 \text{ kJ}}{1 \text{ mol}_{\text{rxn}}} = -486 \text{ kJ}$$

The amount of heat released is 486 kJ.

1 point is earned for a correct calculation (negative sign is not required).

The combustion reaction occurred in a closed room containing $5.56 \times 10^4 \text{ g}$ of air originally at 21.7°C . Assume that all of the heat produced by the reaction was absorbed by the air (specific heat = $1.005 \text{ J/g}^\circ\text{C}$) in the room.

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**Leonard Maltin, Casey St Charnez, Alvin
H Marill, Mike Clark, Cathleen
Anderson, Luke Sader, Bill Warren, Rob
Edelman, Spencer Green**

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