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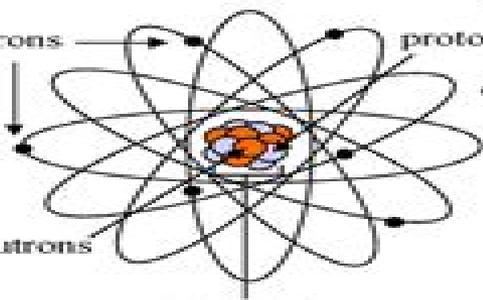
The Atom

Everything is made of **atoms**. Atoms are the smallest part of matter.
Atoms are made up of 3 subatomic particles (particles smaller than the atom): *electrons, protons, and neutrons*.

On the following diagram of an atom define the parts of the atom.

Negatively charged particles; outside the nucleus; can be gained or lost to and from other atoms; very small mass; 1/2000th of a proton

electrons →



protons →

Positively charged particles; in the nucleus; determine what element an atom is.

This model of the atom looks a lot like a solar system. The nucleus, which contain the protons and neutrons, in the center would be the sun. The electrons are the planets spinning around the nucleus.

Neutral particles in the nucleus; give mass to the atom, but not charge.

neutrons →

nucleus

Center of the atom; contains protons and neutrons.

Count the protons to tell what element this is:
of Protons: _____ Element: _____

John Dalton in 1808 published a theory of the atom that had these important points:

- All atoms of a particular element are the same.
- Atoms of different elements have different properties, mass, and chemical reactivity.
- Atoms are not changed by chemical reactions, just rearranged in order or number.

Atoms, Molecules, and Compounds

Atoms combine into **molecules**.
O is an atom; O₂ is a molecule; both are oxygen.

Molecules are made up of two or more atoms.

If two different atoms combine they make **compounds**.
H₂O is a compound; O₂ is a molecule.

Compounds are made up of two or more elements.

Water molecule—
a compound (H₂O)



Oxygen molecule—
an element (O₂)



Atom, molecule or compound?
NaCl — _____
Cl₂ — _____
Na — _____

What elements are these?
Na — _____
Cl — _____

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Léo Pomar



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