

$$dG = -SdT + VdP + \sum_{i=1}^c \mu_i dN_i$$

The Gibbs Equation

$$0 = \sum_{i=1}^c \mu_i dN_i$$

When a system is in Equilibrium, the change in Gibb's energy is 0. In our example, neither the volume nor the temperature of the system change.

$$0 = \sum_{i=1}^c \mu_i dN_i = \mu_1 + \mu_2 + \dots + \mu_c$$

An expansion of the total change in chemical potential for each component with respect to the change in phase.

$$0 = \mu_1 = \mu_1^{(1)} + \mu_1^{(2)} + \dots + \mu_1^{(N_p)}$$

Because our system consists of only one component, the total chemical potential is only that contained in the component (1).

$$0 = \mu_1^{(1)} + \mu_1^{(2)}$$

Because we only have two phases, the summation of the chemical potential of those two phases is equal to 0.

$$\mu_1^{(1)} = -\mu_1^{(2)}$$

This allows us to see that at Equilibrium, the chemical potential of the liquid phase and the vapor phase are equal. The signs are opposite because on change in phase would release the Gibbs Energy that is taken in by the other during a change in phase.

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