

Calculating Average Atomic Mass Worksheet

Teacher

1. The element copper has naturally occurring isotopes with mass numbers of 63 and 65. The relative abundance and atomic masses are 69.2% for a mass of 62.93amu and 30.8% for a mass of 64.93amu. Calculate the average atomic mass of copper.

$$0.692(62.93) + 0.308(64.93) = 63.5 \text{ amu}$$

2. Calculate the average atomic mass of sulfur if 95.00% of all sulfur atoms have a mass of 31.972 amu, 0.76% has a mass of 32.971amu and 4.22% have a mass of 33.967amu.

$$0.9500(31.972) + 0.0076(32.971) + 0.0422(33.967) = 32.1 \text{ amu}$$

3. The four isotopes of lead are shown below, each with its percent by mass abundance and the composition of its nucleus. Using the following data, first calculate the approximate atomic mass of each isotope. Then calculate the average atomic mass of lead.

82p 122n 1.37%	} 204 (^{mass of isotope})	82p 124n 26.26%	82p 125n 20.82%	82p 126n 51.55%
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$$0.0137(204) + 0.2626(206) + 0.2082(207) + 0.5155(208) = 207.2 \text{ amu}$$

4. There are three isotopes of silicon. They have mass numbers of 28, 29 and 30. The average atomic mass of silicon is 28.086amu. What does this say about the relative abundances of the three isotopes?

Since the avg. atomic mass is close to 28, Si-28 must have a much greater percent abundance than Si-29 & Si-30.

5. Calculate the average atomic mass of bromine. One isotope of bromine has an atomic mass of 78.92amu and a relative abundance of 50.69%. The other major isotope of bromine has an atomic mass of 80.92amu and a relative abundance of 49.31%.

$$(0.5069)(78.92) + (0.4931)(80.92) = 79.91 \text{ amu}$$

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S Marginson



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