

ANSWER KEY

Chemistry Semester 3 Study Guide

This document is for your study of the material covered in the following Unit and related associated resources. It contains all answers to the Unit. Be sure to complete all associated resources, practice questions.

Unit 3: Bonding

1. Complete the following table for properties of ionic, covalent, and metallic bonds.

Property	Ionic Bonds	Covalent Bonds	Metallic Bonds
Appearance	Crystalline	Opaque solid liquid or gas at room temperature	Shiny metal/metalloid
Melting Point (High/Low/Variable)	High	Low	Medium
Sublimable solid	Variable	Some are soluble	Not sublimable
Conductivity (as a solid)	No	No	Yes
Conductivity (as a liquid)	Yes	No	Yes
Electrolysis (decomposes by adding)	Yes	No	No
Electron transfer in chemical compounds (yes)	Always + sometimes	Electrovalence	At metals
Electrons in _____ of molecules. (Do not include shared electron bonding in molecules)	Transfer	Sharing	Sharing (free electrons)
Electrolysis (melting/freezing of solids)	Nonmetals producing strong electrolytic solution	None	Electrolysis
Other electrolytic (solvent)	None	None	Yes
Do all the elements form molecules?		 <p>The diagram shows two molecules. The first is a ball-and-stick model of a diatomic molecule with two atoms connected by a bond. The second is a ball-and-stick model of a water molecule (H2O) with one oxygen atom and two hydrogen atoms connected by bonds.</p>	

2. Complete the following table. Answer (properties of compound) below.

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