

Practicing Electron Configuration

1. In the space below, write the **unabbreviated** electron configuration of the following elements.

- i. Sodium _____
- ii. Iron _____
- iii. Bromine _____
- iv. Barium _____
- v. Radium _____

2. In the space below, write the **abbreviated** electron configuration of the following elements.

- i. Cobalt _____
- ii. Silver _____
- iii. Tellurium _____
- iv. Radium _____
- v. Lawrencium _____

3. Determine what elements are denoted by the following electron configuration.

- i. $1s^2 2s^2 2p^6 3s^2 3p^4$: _____
- ii. $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^1$: _____
- iii. $[\text{Kr}] 5s^2 4d^{10} 5p^3$: _____
- iv. $[\text{Xe}] 6s^2 4f^4 5d^6$: _____
- v. $[\text{Rn}] 7s^2 5f^{11}$: _____

4. Answer the following questions

- i. What is the shape of an s orbital? _____
- ii. How many s orbitals can there be in an energy level? _____
- iii. How many electrons can occupy an s orbital? _____
- iv. What is the shape of a p orbital? _____
- v. How many p orbitals can there be in an energy level? _____

Election Configuration Review Worksheet Answers

Mocktime Publication



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