

Reference Tables for Physical Setting/EARTH SCIENCE

Radioactive Decay Data

RADIOACTIVE ISOTOPE	DISINTEGRATION	HALF-LIFE (years)
Carbon-14	$^{14}\text{C} \rightarrow ^{14}\text{N}$	5.7×10^3
Potassium-40	$^{40}\text{K} \rightarrow ^{40}\text{Ar}$ $^{40}\text{K} \rightarrow ^{40}\text{Ca}$	1.3×10^9
Uranium-238	$^{238}\text{U} \rightarrow ^{206}\text{Pb}$	4.5×10^9
Rubidium-87	$^{87}\text{Rb} \rightarrow ^{87}\text{Sr}$	4.9×10^{10}

Specific Heats of Common Materials

MATERIAL	SPECIFIC HEAT (Joules/gram • °C)
Liquid water	4.18
Solid water (ice)	2.11
Water vapor	2.00
Dry air	1.01
Basalt	0.84
Granite	0.79
Iron	0.45
Copper	0.38
Lead	0.13

Equations

$$\text{Eccentricity} = \frac{\text{distance between foci}}{\text{length of major axis}}$$

$$\text{Gradient} = \frac{\text{change in field value}}{\text{distance}}$$

$$\text{Rate of change} = \frac{\text{change in value}}{\text{time}}$$

$$\text{Density} = \frac{\text{mass}}{\text{volume}}$$

Properties of Water

Heat energy gained during melting	334 J/g
Heat energy released during freezing	334 J/g
Heat energy gained during vaporization	2260 J/g
Heat energy released during condensation	2260 J/g
Density at 3.98°C	1.0 g/mL

Average Chemical Composition of Earth's Crust, Hydrosphere, and Troposphere

ELEMENT (symbol)	CRUST		HYDROSPHERE	TROPOSPHERE
	Percent by mass	Percent by volume	Percent by volume	Percent by volume
Oxygen (O)	46.10	94.04	33.0	21.0
Silicon (Si)	28.20	0.88		
Aluminum (Al)	8.23	0.48		
Iron (Fe)	5.63	0.49		
Calcium (Ca)	4.15	1.18		
Sodium (Na)	2.36	1.11		
Magnesium (Mg)	2.33	0.33		
Potassium (K)	2.09	1.42		
Nitrogen (N)				78.0
Hydrogen (H)			66.0	
Other	0.91	0.07	1.0	1.0

2011 EDITION

This edition of the Earth Science Reference Tables should be used in the classroom beginning in the 2011–12 school year. The first examination for which these tables will be used is the January 2012 Regents Examination in Physical Setting/Earth Science.



Earth Science Reference Table Workbook 3rd Edition

Dionysius Lardner



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Table of Contents Earth Science Reference Table Workbook 3rd Edition

1. Understanding the eBook Earth Science Reference Table Workbook 3rd Edition
 - The Rise of Digital Reading Earth Science Reference Table Workbook 3rd Edition
 - Advantages of eBooks Over Traditional Books
2. Identifying Earth Science Reference Table Workbook 3rd Edition
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Earth Science Reference Table Workbook 3rd Edition
 - User-Friendly Interface
4. Exploring eBook Recommendations from Earth Science Reference Table Workbook 3rd Edition
 - Personalized Recommendations
 - Earth Science Reference Table Workbook 3rd Edition User Reviews and Ratings

- Earth Science Reference Table Workbook 3rd Edition and Bestseller Lists
- 5. Accessing Earth Science Reference Table Workbook 3rd Edition Free and Paid eBooks
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 - Earth Science Reference Table Workbook 3rd Edition Budget-Friendly Options
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 - ePub, PDF, MOBI, and More
 - Earth Science Reference Table Workbook 3rd Edition Compatibility with Devices
 - Earth Science Reference Table Workbook 3rd Edition Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Earth Science Reference Table Workbook 3rd Edition
 - Highlighting and Note-Taking Earth Science Reference Table Workbook 3rd Edition
 - Interactive Elements Earth Science Reference Table Workbook 3rd Edition
- 8. Staying Engaged with Earth Science Reference Table Workbook 3rd Edition
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Earth Science Reference Table Workbook 3rd Edition
- 9. Balancing eBooks and Physical Books Earth Science Reference Table Workbook 3rd Edition
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Earth Science Reference Table Workbook 3rd Edition
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Earth Science Reference Table Workbook 3rd Edition
 - Setting Reading Goals Earth Science Reference Table Workbook 3rd Edition
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Earth Science Reference Table Workbook 3rd Edition
 - Fact-Checking eBook Content of Earth Science Reference Table Workbook 3rd Edition
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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