

1.1 Chemical Reactions

1.1.1 Describing Chemical Reactions

Read to your text to understand how to write chemical equations. Remember that you will be using mass balancing the equations to get the best number of atoms.

1. Write the equation for the combustion of methane and hydrogen sulfide respectively:



3. Write the equation for the decomposition of hydrogen peroxide:



4. Write the equation for the reaction of sodium metal with chlorine gas to form sodium chloride:



1.1.1.1 Balancing Chemical Equations

1. In chemical reactions matter is not created or destroyed. Change this into an equation: $\text{mass}_{\text{reactants}} = \text{mass}_{\text{products}}$

2. The sum of coefficients is not equal to the sum of subscripts. $2 + 1 \neq 2 + 1$

3. In the following reactions, the atoms are balanced, but the charges are not balanced. In general, they are balancing the atoms. $\text{Fe} + \text{O}_2 \rightarrow \text{Fe}_2\text{O}_3$

4. Write the following reactions that describe the following processes:

- a. Ferrous sulfate reacts with oxygen to form ferric sulfate and sulfur dioxide.
- b. Ferrous sulfate decomposes to form ferric sulfate and sulfur dioxide.

5. Write a balanced equation:

6. In chemical reactions the atoms are not created or destroyed but are the atoms themselves that remain independent in time. $\text{Fe} + \text{O}_2 \rightarrow \text{Fe}_2\text{O}_3$

7. Identify the reactants and products of hydrochloric acid and sodium hydroxide:

a. reactants: HCl and NaOH b. products: NaCl and H_2O



Answer Key To Section 11 1 Describing Chemical Reactions

Emilie Sanchez



Answer Key To Section 11 1 Describing Chemical Reactions:

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