

## Empirical and Molecular Formula Worksheet ANSWER KEY

Write the empirical formula for the following compounds.

- 1)  $C_6H_6$     **CH**
- 6)  $C_8H_{12}$      **$C_4H_6$**
- 7)  $WO_2$      **$WO_2$**
- 8)  $C_2H_6O_2$      **$CH_3O$**
- 9)  $X_{10}Y_{12}$      **$X_5Y_6$**
- 6) A compound with an empirical formula of  $C_2OH_4$  and a molar mass of 88 grams per mole. What is the molecular formula of this compound?     **$C_8O_4H_{16}$**  
- 7) A compound with an empirical formula of  $C_4H_6O$  and a molar mass of 136 grams per mole. What is the molecular formula of this compound?     **$C_8H_{12}O_2$**  
- 8) A compound with an empirical formula of  $CFBrO$  and a molar mass of 253.8 grams per mole. What is the molecular formula of this compound?     **$C_2F_2Br_2O_2$**  
- 9) A compound with an empirical formula of  $C_2H_3N$  and a molar mass of 46 grams per mole. What is the molecular formula of this compound?     **$C_2H_3N$**  
- 10) A well-known reagent in analytical chemistry, dimethylglyoxime, has the empirical formula  $C_2H_5NO$ . If its molar mass is 116.1 g/mol, what is the molecular formula of the compound?     **$C_4H_{10}N_2O_2$**  
12. A certain blue solid contains 36.84% N and 63.16% O. What is the empirical formula of this compound?    **The ratios are  $N_{1.50}O_{1.50}$ . Since 1.50 is not close to a whole number, we multiply *both* subscripts by 2. The empirical formula is thus  $N_2O_2$ . (The name is dinitrogen trioxide.)** 
13. A sample of indium chloride weighing 0.5000 g is found to contain 0.2404 g of chlorine. What is the empirical formula of the indium compound?     **$InCl_2$**  
14. An unknown compound was found to have a percent composition as follows: 47.0 % potassium, 14.5 % carbon, and 38.5 % oxygen. What is its empirical formula? If the true molar mass of the compound is 166.22 g/mol, what is its molecular formula?     **$K_2C_3O_4$**  
15. Rubbing alcohol was found to contain 60.0 % carbon, 13.4 % hydrogen, and the remaining mass was due to oxygen. What is the empirical formula of rubbing alcohol?     **$C_3H_8O$**  

# Answer Key To Chemistry A Molecular Approach

**Siddharth Venkatesh**



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Jean-Marie Basset, Bruce C. Gates, Jean-Pierre Candy, Agnès Choplin, Michel Leconte, Françoise Quignard, Cathérine Santini, 2012-12-06 Surface organometallic chemistry is a new field bringing together researchers from organometallic inorganic and surface chemistry and catalysis Topics ranging from reaction mechanisms to catalyst preparation are considered from a molecular basis according to which the active site on a catalyst surface has a supra molecular character This the first book on the subject is the outcome of a NATO Workshop held in Le Rouret France in May 1986 It is our hope that the following chapters and the concluding summary of recommendations for research may help to provide a definition of surface organometallic chemistry Besides catalysis the central theme of the Workshop four main topics are considered 1 Reactions of organometallics with surfaces of metal oxides metals and zeolites 2 Molecular models of surfaces metal oxides and metals 3 Molecular approaches to the mechanisms of surface reactions 4 Synthesis and modification of zeolites and related microporous solids Most surface organometallic chemistry has been carried out on amorphous high surface area metal oxides such as silica alumina magnesia and titania The first chapter contributed by KNOZINGER gives a short summary of the structure and reactivity of metal oxide surfaces Most of our understanding of these surfaces is based on acid base and redox chemistry this chemistry has developed from X ray and spectroscopic data and much has been inferred from the structures and reactivities of adsorbed organic probe molecules There are major opportunities for extending this understanding by use of well defined single crystal oxide surfaces and organometallic probe molecules

**Chemistry**

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