

1. As a solute is added to a solvent, what happens to the freezing point and the boiling point of the solution?
- A) The freezing point decreases and the boiling point increases.
- B) The freezing point increases and the boiling point increases.
- C) The freezing point increases and the boiling point decreases.
- D) The freezing point decreases and the boiling point decreases.
2. Compared to the freezing point of 1.0 M KCl(aq) at standard pressure, the freezing point of 1.0 M CaCl<sub>2</sub>(aq) at standard pressure is
- More ions*  A) lower
- B) higher
- C) the same
3. Which aqueous solution of KI freezes at the lowest temperature?
- A) 1 mol of KI in 500. g of water
- B) 2 mol of KI in 500. g of water *(more concentrated)*
- C) 1 mol of KI in 1000. g of water
- D) 2 mol of KI in 1000. g of water
4. A 1 kilogram sample of water will have the highest freezing point when it contains
- A)  $1 \times 10^{21}$  dissolved particles
- B)  $1 \times 10^{17}$  dissolved particles
- C)  $1 \times 10^{19}$  dissolved particles
- D)  $1 \times 10^{23}$  dissolved particles *lowest f.p.*
5. A solution consists of 0.50 mole of CaCl<sub>2</sub> dissolved in 100. grams of H<sub>2</sub>O at 25°C. Compared to the boiling point and freezing point of 100. grams of H<sub>2</sub>O at standard pressure, the solution at standard pressure has
- A) a higher boiling point and a lower freezing point
- B) a lower boiling point and a lower freezing point
- C) a higher boiling point and a higher freezing point
- D) a lower boiling point and a higher freezing point
6. Which solution will freeze at the lowest temperature?
- A) 1 mole of sugar in 500 g of water
- B) 1 mole of sugar in 1,000 g of water
- C) 2 moles of sugar in 1,000 g of water
- D) 2 moles of sugar in 500 g of water
7. Which solution has the highest boiling point at standard pressure?
- A) 0.10 M K<sub>3</sub>PO<sub>4</sub>(aq) *4 ions 3 K<sup>+</sup> 3 PO<sub>4</sub><sup>-</sup>*
- B) 0.10 M KCl(aq)
- C) 0.10 M KNO<sub>3</sub>(aq)
- D) 0.10 M K<sub>2</sub>SO<sub>4</sub>(aq)
8. Which concentration of a solution of CH<sub>3</sub>OH in water has the lowest freezing point?
- A) 0.001 M
- B) 0.1 M
- C) 0.01 M
- D) 0.0001 M
9. How do the boiling point and freezing point of a solution of water and calcium chloride at standard pressure compare to the boiling point and freezing point of water at standard pressure?
- A) Both the freezing point and boiling point of the solution are higher.
- B) The freezing point of the solution is higher and the boiling point of the solution is lower.
- C) The freezing point of the solution is lower and the boiling point of the solution is higher.
- D) Both the freezing point and boiling point of the solution are lower.
10. Compared to a 2.0 M aqueous solution of NaCl at 1 atmosphere, a 3.0 M aqueous solution of NaCl at 1 atmosphere has a
- A) lower boiling point and a higher freezing point
- B) higher boiling point and a higher freezing point
- C) lower boiling point and a lower freezing point
- D) higher boiling point and a lower freezing point

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*manual 118 hartzell overhaul aeroval - Mar 15 2022*

web manual 118 nomenclature hartzell overhaul supersedure reference manufactuer oem hawker beechcraft textron location wichita ks usa website txtav com cage code fscm 70898 product classification aircraft part model reference beech commercial spares market segment aftermarket aircraft type

*hartzell hc y 1 series owner s manual and logbook - Feb 23 2023*

web page 17 federal aviation administration rev description of revision adds airworthiness limitation information from hartzell overhaul manual 113b 61 10 13 and hartzell overhaul manual 117d 61 10 17 added cylinder life limits for propeller model hc c 2 3 yr 4 b c f fc8477 4r and corrected the engine model number for aviatt pitts s