

**CHAPTER 4 PRACTICE TEST: SOLUTION STOICHIOMETRY**

1) Which of the following are weak electrolytes? HCl,  $\text{HC}_2\text{H}_3\text{O}_2$ ,  $\text{NH}_3$ , KCl

A) HCl, KCl

B) HCl,  $\text{HC}_2\text{H}_3\text{O}_2$ ,  $\text{NH}_3$ , KCl

C)  $\text{HC}_2\text{H}_3\text{O}_2$ , KCl

D)  $\text{HC}_2\text{H}_3\text{O}_2$ ,  $\text{NH}_3$

E) HCl,  $\text{HC}_2\text{H}_3\text{O}_2$ , KCl

Answer: D Diff: 2

2) What are the spectator ions in the reaction between KOH (aq) and  $\text{HNO}_3$ (aq)?

A)  $\text{K}^+$  and  $\text{H}^+$

C)  $\text{K}^+$  and  $\text{NO}_3^-$

E)  $\text{OH}^-$  only

B)  $\text{H}^+$  and  $\text{OH}^-$

D)  $\text{H}^+$  and  $\text{NO}_3^-$

Answer: C Diff: 2

3) The balanced net ionic equation for precipitation of  $\text{CaCO}_3$  when aqueous solutions of  $\text{Na}_2\text{CO}_3$  and  $\text{CaCl}_2$  are mixed is \_\_\_\_\_.

A)  $2\text{Na}^+(\text{aq}) + \text{CO}_3^{2-}(\text{aq}) \rightarrow \text{Na}_2\text{CO}_3(\text{aq})$

E)

B)  $2\text{Na}^+(\text{aq}) + 2\text{Cl}^-(\text{aq}) \rightarrow 2\text{NaCl}(\text{aq})$

$\text{Na}_2\text{CO}_3(\text{aq}) + \text{CaCl}_2(\text{aq}) \rightarrow 2\text{NaCl}(\text{aq}) + \text{CaCO}_3(\text{s})$

C)  $\text{Na}^+(\text{aq}) + \text{Cl}^-(\text{aq}) \rightarrow \text{NaCl}(\text{aq})$

D)  $\text{Ca}^{2+}(\text{aq}) + \text{CO}_3^{2-}(\text{aq}) \rightarrow \text{CaCO}_3(\text{s})$

Answer: D Diff: 3

4) Which of the following are strong acids? HI,  $\text{HNO}_3$ , HF, HBr

A) HF, HBr

C) HI, HF, HBr

E) HI,  $\text{HNO}_3$ , HBr

B) HI,  $\text{HNO}_3$ , HF, HBr

D)  $\text{HNO}_3$ , HF, HBr

Answer: E Diff: 2

5) A neutralization reaction between an acid and a metal hydroxide produces \_\_\_\_\_.

A) water and a salt

C) oxygen gas

E) ammonia

B) hydrogen gas

D) sodium hydroxide

Answer: A Diff: 2

6) Which of these metals is the least easily oxidized? Na, Au, Fe, Ca, Ag

A) Na

B) Au

C) Fe

D) Ca

E) Ag

Answer: B Diff: 3

7) Based on the equations below, which metal is the most active?



A) Ni

B) Ag

C) Cu

D) Pb

E) N

Answer: A Diff: 3

8) What is the concentration (M) of KCl in a solution made by mixing 25.0 mL of 0.100 M KCl with 50.0 mL of 0.100 M KCl?

A) 0.100

B) 0.0500

C) 0.0333

D) 0.0250

E) 125

Answer: A Diff: 3

# Answer Key A Natural Approach To Chemistry

**Harry Clary Jones**



## **Answer Key A Natural Approach To Chemistry:**

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communities which vary depending on their baseline vulnerability These include risks to human life human health property  
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life under these changing conditions Some examples of adaptation include integrating robust and flexible operations  
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ecosystem or community based approaches strengthening stakeholders for decision making designing new financing  
mechanisms and using data driven methods to improve resilience One important area of focus for climate adaptation is  
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processes into various systems The drivers for integration can include increasing efficiency improving environmental impacts  
creating multiple benefits and reducing resource use to name a few NbS activities have been used to upgrade water systems  
reduce heat island effects improve stormwater management green infrastructure building and upgrading efforts and provide  
multiple benefits to both ecosystems and communities While NbS is not a novel concept more research on the conceptual and  
technical aspects is needed to form a strong foundation from which applications can be designed Research on the potential of  
NbS activities for specific climate change adaptation needs could unveil key insights in applicability measurement  
effectiveness and scalability The types of research needed to address climate change adaptation through NbS are very wide  
ranging The presentation of conceptual frameworks and links between NbS and similar concepts including ecosystem  
services ecosystem based adaptation green infrastructure and passive treatment is important to articulate and clarify for use  
and integration in the climate adaptation context Evaluating the targeted implementation of NbS across landscapes coastal  
regions and watersheds as well as exploring its applicability in sectors like energy water forestry agriculture industry and the  
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