

# Chapter 21 -- Nuclear Chemistry

## 21.1 Radioactivity

- When nuclei change spontaneously, emitting energy, they are said to be **radioactive**.
- **Nuclear chemistry** is the study of nuclear reactions and their uses.
- **Isotopes** have the same number of  $p^+$  but different numbers of  $n^0$ .
  - Different isotopes of the same element are distinguished by their mass numbers.
  - Different isotopes have different natural abundances.
- A **radionuclide** is a radioactive nucleus.
  - Atoms containing these nuclei are called **radioisotopes**.

### Nuclear Equations

- Most nuclei are stable.
  - Radionuclides are unstable and spontaneously emit particles and/or electromagnetic radiation.
  - Example: Uranium-238 is radioactive.
    - It emits **alpha ( $\alpha$ ) particles** (helium-4 nuclei)
- When a nucleus spontaneously decomposes in this manner, we say it has *decayed (radioactive decay)*.
- In nuclear equations, the total number of nucleons is conserved.
  - We can represent the uranium-238 decay by the following nuclear equation:



- The total number of protons and neutrons before a nuclear reaction must be the same as the total number of nucleons after the reaction.

### Types of Radioactive Decay

- There are three types of radiation which we will consider:
  - $\alpha$ -Radiation is the loss of  ${}_2^4\text{He}$  (**alpha particles**) from the nucleus.
  - $\beta$ -Radiation is the loss of an electron from the nucleus. These high-speed electrons are called **beta particles**.
  - $\gamma$ -Radiation is the loss of high-energy photons from the nucleus.
- Nucleons can undergo two other types of decay:
  - **positron emission**
    - A positron is a particle with the same mass as an electron but with the opposite sign.
  - **electron capture**
    - The nucleus captures an electron from the electron cloud surrounding the nucleus.
- Representations:
  - In nuclear chemistry, to ensure the conservation of nucleons we write all particles with their atomic and mass numbers:  ${}_2^4\text{He}$  and  ${}_{-1}^0\text{e}$  represent  $\alpha$ - and  $\beta^-$  radiation.
- Nucleons can undergo decay:
  - ( $\beta^-$ -emission)  ${}_0^1\text{n} \rightarrow {}_1^1\text{p} + {}_{-1}^0\text{e}$
  - (positron annihilation)  ${}_{-1}^0\text{e} + {}_1^0\text{e} \rightarrow 2{}_0^0\gamma$
  - (positron or  $\beta^+$ -emission)  ${}_1^1\text{p} \rightarrow {}_0^1\text{n} + {}_1^0\text{e}$
  - (electron capture)  ${}_1^1\text{p} + {}_{-1}^0\text{e} \rightarrow {}_0^1\text{n}$

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