

## 20

**ACIDS AND BASES**  
PRACTICE PROBLEMS

In your notebook, solve the following problems.

**SECTION 20.1 DESCRIBING ACIDS AND BASES**

1. Name the following compounds as acids.
  - a.  $\text{HNO}_2$
  - b.  $\text{H}_2\text{SO}_4$
  - c. HF
  - d.  $\text{H}_2\text{CO}_3$
2. Write the formulas for the following bases.
  - a. calcium hydroxide
  - b. ammonium hydroxide
  - c. aluminum hydroxide
  - d. lithium hydroxide
3. Describe some distinctive properties of acids.
4. Describe some distinctive properties of bases.

**SECTION 20.2 HYDROGEN IONS AND ACIDITY**

1. A solution has a hydrogen ion concentration of  $1 \times 10^{-6}\text{M}$ . What is its pH?
2. What is the pH of a solution if the  $[\text{H}^+] = 7.2 \times 10^{-9}\text{M}$ ?
3. What is the pOH of a solution if the  $[\text{OH}^-] = 3.5 \times 10^{-3}\text{M}$ ?
4. What is the pOH of a solution that has a pH of 3.4?
5. Classify each solution as acidic, basic, or neutral.
  - a.  $[\text{H}^+] = 2.5 \times 10^{-9}\text{M}$
  - b.  $\text{pOH} = 12.0$
  - c.  $[\text{OH}^-] = 9.8 \times 10^{-11}\text{M}$
  - d.  $[\text{H}^+] = 1 \times 10^{-7}\text{M}$
  - e.  $\text{pH} = 0.8$
6. Calculate the pH of each solution.
  - a.  $[\text{H}^+] = 1 \times 10^{-5}\text{M}$
  - b.  $[\text{H}^+] = 4.4 \times 10^{-12}\text{M}$
  - c.  $[\text{OH}^-] = 2.2 \times 10^{-7}\text{M}$
  - d.  $\text{pOH} = 1.4$
7. Classify the solutions in problem 6 as acidic or basic.
8. Why is there a minus sign in the definition of pH?
9. A solution has a pOH of 12.4. What is the pH of this solution?
10. What is the pH of a solution with  $[\text{H}^+] = 1 \times 10^{-3}\text{M}$ ?

**SECTION 20.3 ACID-BASE THEORIES**

1. Identify the hydrogen ion donor(s) and hydrogen ion acceptor(s) for ionization of  $\text{H}_2\text{SO}_4$  in water. Label the conjugate acid-base pairs.
2. Identify all of the ions that may be formed when  $\text{H}_3\text{PO}_4$  ionizes in water.

# Chapter 20 Acids Bases Practice Problems Answers

**Julie R. Frentrup**



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