

Name \_\_\_\_\_ Date \_\_\_\_\_ Class \_\_\_\_\_

## CHAPTER

## 15

## STUDY GUIDE FOR CONTENT MASTERY

## Solutions

### Section 15.1 What are solutions?

In your textbook, read about the characteristics of solutions.

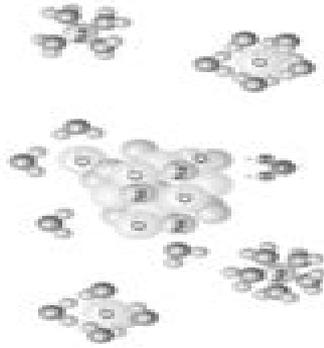
Use each of the terms below just once to complete the passage.

immiscible	liquid	soluble	solution
insoluble	miscible	solvent	solvent

Air is a(n) **(1)** solution of oxygen gas dissolved in nitrogen gas. The oxygen is an **(2)** solute, and nitrogen is the **(3)** solvent. Because oxygen gas dissolves in a solvent, oxygen gas is an **(4)** soluble substance. A substance that does not dissolve is **(5)** insoluble. **(6)** Liquid solutions are the most common type of solutions. If one liquid is soluble in another liquid, such as acetic acid in water, the two liquids are **(7)** miscible. However, if one liquid is insoluble in another, the liquids are **(8)** immiscible.

Read about hydration in aqueous solutions in your textbook.

The diagram shows the hydration of solid sodium chloride to form an aqueous solution. Use the diagram to answer the following questions.



9. Hydration is hydration in which the solvent is water. What is hydration?  
**Solvation is the process of surrounding solute particles with solvent particles to form a solution.**

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## STUDY GUIDE FOR CONTENT MASTERY

## CHAPTER

## 15

## Section 15.1 continued

16. An sodium chloride dissolves in water. What happens to the sodium and chloride ions? **The sodium and chloride ions are separated and surrounded by the water molecules.**

17. Explain the orientation of the water molecules around the sodium ions and chloride ions. **Because the sodium ion is positively charged, it attracts the negatively charged portion of the water molecule (the oxygen atom) and repels the positively charged portion of the water molecule (the hydrogen atoms). Because the chloride ion is negatively charged, it attracts the positively charged portion of the water molecule and repels the negatively charged portion.**

18. How does the strength of the attraction between water molecules and sodium and chloride ions compare with the strength of the attraction between the sodium ions and chloride ions? **How do you know?**

**The attraction between the water molecules and the sodium and chloride ions is greater than the attraction between the sodium and chloride ions. The greater strength of attraction between the water molecules and the ions is what causes the solvation process to occur.**

19. List three ways that the rate of solvation may be increased.

**stirring or shaking the solution, breaking the solute into smaller pieces, and heating the solvent**

In your textbook, read about heat of solution, solubility, and factors that affect solubility.

For each statement below, write *true* or *false*.

- true** 14. The overall energy change that occurs when a solution forms is called the heat of solution.
- false** 15. Solubility is a measure of the maximum amount of solute that dissolves in a given amount of solvent at a specified temperature and pressure.
- false** 16. Solvation continues as long as the solvation rate is less than the crystallization rate.
- true** 17. In a saturated solution, solvation and crystallization are in equilibrium.
- true** 18. Additional water can be dissolved in an unsaturated solution.
- true** 19. The solubility of a gas dissolved in a liquid decreases as the temperature of the solution increases.

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# Chapter 1stoichiometry Study Guide For Content Mastery Key

**Glencoe/McGraw-Hill**



## **Chapter 1stoichiometry Study Guide For Content Mastery Key:**

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In a period characterized by interconnectedness and an insatiable thirst for knowledge, the captivating potential of verbal expression has emerged as a formidable force. Its power to evoke sentiments, stimulate introspection, and incite profound transformations is genuinely awe-inspiring. Within the pages of "**Chapter 1stoichiometry Study Guide For Content Mastery Key**," a mesmerizing literary creation penned by a celebrated wordsmith, readers set about an enlightening odyssey, unraveling the intricate significance of language and its enduring impact on our lives. In this appraisal, we shall explore the book's central themes, evaluate its distinctive writing style, and gauge its pervasive influence on the hearts and minds of its readership.

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