

## Chapter 21 Physical properties of gases

### Q1.

Use the kinetic molecular theory to explain the following observed properties of gases:

- a Gases occupy all the available space in a container.
- b Gases can be easily compressed compared with their corresponding liquid forms.
- c A given volume of a gaseous substance weighs less than the same volume of the substance in the liquid state.
- d Gases will mix together readily.
- e The total pressure of a mixture of gases is equal to the sum of the pressures exerted by each of the gases in the mixture.

### A1.

- a Molecules of gases are in constant, rapid, random motion and the forces between molecules are negligible. They continue to move outwards until stopped by the walls of the container, filling all the space available.
- b Most of the volume occupied by a gas is space, so compression can be achieved by reducing the space between the particles.
- c The molecules in a gas are spread much further apart than those of a liquid. A given mass of gas would occupy a much greater volume than the same mass of the liquid phase. Therefore, the density of the gas is less.
- d Gases mix easily together because of the large amount of space between the molecules.
- e The pressure exerted by a gas depends on the number of collisions of gas particles and the wall of the container. The pressure is independent of the type of gas involved. The total pressure exerted by a mixture of gases will depend on the total number of collisions each gas has with the container.

### Q2.

Use the ideas of the kinetic molecular theory of gases to explain the following observations:

- a Tyre manufacturers recommend a maximum pressure for tyres.
- b The pressure in a car's tyres will increase if a long distance is travelled on a hot day.
- c You can smell dinner cooking as you enter the house.
- d A balloon will burst if you blow it up too much.

### A2.

- a Tyres have a recommended maximum pressure to give a comfortable ride as well as good traction on the road. If the pressure in a tyre is too high, the gas inside cannot be compressed as easily and passengers will be more aware of bumps on the road.
- b During a long journey on a hot day, the air in a tyre warms up. This means the air molecules have increased kinetic energy, and collisions with the walls of the tyres will increase, and so the pressure will increase.

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**Marylou Morano-Kjelle**



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## **Table of Contents Chapter 21 Physical Properties Of Gases Wikispaces**

1. Understanding the eBook Chapter 21 Physical Properties Of Gases Wikispaces
  - The Rise of Digital Reading Chapter 21 Physical Properties Of Gases Wikispaces
  - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 21 Physical Properties Of Gases Wikispaces
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chapter 21 Physical Properties Of Gases Wikispaces
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 21 Physical Properties Of Gases Wikispaces
  - Personalized Recommendations
  - Chapter 21 Physical Properties Of Gases Wikispaces User Reviews and Ratings
  - Chapter 21 Physical Properties Of Gases Wikispaces and Bestseller Lists

5. Accessing Chapter 21 Physical Properties Of Gases Wikispaces Free and Paid eBooks
  - Chapter 21 Physical Properties Of Gases Wikispaces Public Domain eBooks
  - Chapter 21 Physical Properties Of Gases Wikispaces eBook Subscription Services
  - Chapter 21 Physical Properties Of Gases Wikispaces Budget-Friendly Options
6. Navigating Chapter 21 Physical Properties Of Gases Wikispaces eBook Formats
  - ePub, PDF, MOBI, and More
  - Chapter 21 Physical Properties Of Gases Wikispaces Compatibility with Devices
  - Chapter 21 Physical Properties Of Gases Wikispaces Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chapter 21 Physical Properties Of Gases Wikispaces
  - Highlighting and Note-Taking Chapter 21 Physical Properties Of Gases Wikispaces
  - Interactive Elements Chapter 21 Physical Properties Of Gases Wikispaces
8. Staying Engaged with Chapter 21 Physical Properties Of Gases Wikispaces
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chapter 21 Physical Properties Of Gases Wikispaces
9. Balancing eBooks and Physical Books Chapter 21 Physical Properties Of Gases Wikispaces
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chapter 21 Physical Properties Of Gases Wikispaces
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Chapter 21 Physical Properties Of Gases Wikispaces
  - Setting Reading Goals Chapter 21 Physical Properties Of Gases Wikispaces
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chapter 21 Physical Properties Of Gases Wikispaces
  - Fact-Checking eBook Content of Chapter 21 Physical Properties Of Gases Wikispaces
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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