

MOLE CONCEPT

CLASS Xth

1. Calculate the relative molecular masses (or molecular weights) of the following compounds :

- i) Copper sulphate crystals, $\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$.
- ii) Ammonium sulphate, $(\text{NH}_4)_2\text{SO}_4$.
- iii) Cane sugar, $\text{C}_{12}\text{H}_{22}\text{O}_{11}$.

(Given atomic masses Cu = 63.5, S = 32, O = 16, N = 14, C = 12)

2. An organic compound with vapour density = 94 contains, C = 12.67%, H = 2.13%, and Br = 85.11%. Find the molecular formula. [Atomic mass : C = 12, H = 1, Br = 80]

3. (i) The compound A has the following percentage composition by mass. Carbon = 26.7%, oxygen = 71.1%, hydrogen = 2.2%. Determine the Empirical formula of compound A [work to one decimal place] [H = 1, C = 12, O = 16].

ii) If the relative molecular mass of A is 90, what is the molecular formula of A?

(iii) The compound A is weak acid. What is meant by this statement?

4. Determine the empirical formula of the compound whose composition by mass is: 42% nitrogen, 48% oxygen and 9% hydrogen. [H = 1; N = 14; O = 16]

5. Determine the empirical formula of a compound containing 47.9% potassium, 5.5% beryllium and 46.6% fluorine by mass. (Atomic weight of Be = 9; F = 19; K = 39) Work to one decimal place.

6. The Empirical formula of a compound is C_2H_5 . It has a vapour density of 29. Determine the relative molecular formula mass of the compound and hence its molecular formula.

7. The percentage composition of sodium phosphate as determined by analysis, is 42.1% sodium, 18.9% phosphorus and 39% oxygen. Find the empirical formula of the compound (work to two decimal places).

8. A gaseous organic compound contains 3.6 g of carbon and 0.8 g of hydrogen. The vapour density of this compound is 22. (i) Calculate the Empirical formula. (ii) Calculate the molecular formula of the compound.

Chapter 1the Mole Answers

**Donald J. Wink, Sharon Fetzer-
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Chapter 1the Mole Answers:

A Textbook of Organic Chemistry, 22e Arun Bahl & B S Bahl, 2019 With an increased focus on fundamentals this new edition of A Textbook of Organic Chemistry continues to present the time tested functional group approach to the subject This examination oriented book breaks the intricacies of Organic Chemistry into easy to understand steps which gives the student the necessary foundation to build upon learn and understand Organic Chemistry in a way that is efficient as well as long lasting Basic Principles of Inorganic Chemistry Brian J Hathaway, Clair Murphy, Brian Murphy, 2007-10-31 General chemistry textbooks are usually lengthy and present chemistry to the student as an unconnected list of facts In inorganic chemistry emphasis should be placed on the connections between valence shell electron configuration and the physical and chemical properties of the element Basic Principles of Inorganic Chemistry Making the Connections is a short concise book that emphasises these connections in particular the chemistry of the Main Group compounds With reference to chemical properties Lewis Structures stoichiometry and spider diagrams students will be able to predict or calculate the chemistry of simple polyatomic compounds from the valence shell configuration and will no longer be required to memorise vast amounts of factual chemistry This book is ideal for students taking chemistry as a subsidiary subject as well as honours degree students **Oswaal JEE (Main) Question Bank Chemistry | Chapter-wise & Topic-wise Solved Papers | 2019-2024 | For 2025 Exam** Oswaal Editorial Board, 2024-02-28 Description of the Product 100% Updated with 2 latest solved papers of 27th January Shift 1 29th January Shift 2 2024 Extensive Practice with more than 1500 fully solved questions of 2019 to 2023 Concept Clarity with Chapter wise Topic wise Concept based videos Mind Maps Mnemonics Valuable Exam Insights with Tips to crack JEE Main Exam in first Attempt Examination Analysis with last 5 Years Chapter wise Trend Analysis **The Practice of Chemistry** Donald J. Wink, Sharon Fetzer-Gislason, Sheila McNicholas, 2003-03 Students can't do chemistry if they can't do the math The Practice of Chemistry First Edition is the only preparatory chemistry text to offer students targeted consistent mathematical support to make sure they understand how to use math especially algebra in chemical problem solving The book's unique focus on actual chemical practice extensive study tools and integrated media makes The Practice of Chemistry the most effective way to prepare students for the standard general chemistry course and bright futures as science majors This special PowerPoint tour of the text was created by Don Wink http://www.bfwpub.com/pdfs/wink/POCPowerPoint_Final.ppt 832KB BRS Biochemistry, Molecular Biology, and Genetics Michael A. Lieberman, Emma C. Schoch, 2024-12-18 Practical approachable and perfect for today's busy medical students and practitioners BRS Biochemistry Molecular Biology and Genetics 8th Edition provides a comprehensive yet concise review to help ensure excellence on class exams and the USMLE Step 1 The renowned Board Review Series outline format offers efficient content review enhanced by key learning aids clinical correlations and USMLE style questions to reinforce understanding and test preparation Aligned with recent changes in medical education and board exams this updated edition emphasizes clinically relevant information to

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Programmed Problem-solving for First-year Chemistry Arnold B. Loebel, 1983 Thermodynamics and Chemistry Frank Henry MacDougall, 1921 Basic Mathematics for Beginning Chemistry Dorothy M. Goldish, 1990 *Chemistry Workbook For Dummies* Peter J. Mikulecky, Katherine Brutlag, Michelle Rose Gilman, Brian Peterson, 2008-08-06 From liquids and solids to acids and bases work chemistry equations and use formulas with ease Got a grasp on the chemistry terms and concepts you need to know but get lost halfway through a problem or worse yet not know where to begin Have no fear this hands on guide helps you solve many types of chemistry problems in a focused step by step manner With problem solving shortcuts and lots of practice exercises you'll build your chemistry skills and improve your performance both in and out of the science lab You'll see how to work with numbers atoms and elements make and remake compounds understand changes in terms of energy make sense of organic chemistry and more 100s of Problems Know where to begin and how to solve the most

common chemistry problems Step by step answer sets clearly identify where you went wrong or right with a problem
Understand the key exceptions to chemistry rules Use chemistry in practical applications with confidence Physical
Science - Chemistry Split with Online Learning Center Password Card (Chapters 1 And 8 - 13) Bill W. Tillery,Tillery
Bill,2004-01-09

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