

(b) At 0°C , the water is at its freezing point, hence both solid and liquid phases are observed.

(c) At 100°C , the water is at its boiling point, hence both liquid and gaseous states of water (water vapour) are observed.

6. Give two reasons to justify –

(a) Water at room temperature is a liquid.

(b) An iron almirah is a solid at room temperature.

Solution:

(a) Water persists as a liquid at room temperature since its melting point is lower than room temperature and its boiling point (100°C) is higher.

Similarly,

(i). A fixed volume is occupied by a fixed mass of water.

(ii). At room temperature, water does not have a fixed shape and flows to fit the container's shape.

As a result, water is a liquid at room temperature.

(b) Because its melting and boiling points are above room temperature, an iron almirah is a solid at room temperature. In the same way,

(i) An iron almirah is rigid and has a predetermined shape.

(ii) Metals have a relatively high density.

As a result, at room temperature, iron almirah is a solid.

7. Why is ice at 273K more effective in cooling than water at the same temperature?

Solution:

At 273K , ice will absorb heat energy or latent heat from the medium to overcome fusion and transform into water. As a result, ice has a greater cooling impact than water at the same temperature since water does not absorb the excess heat from the medium.

8. What produces more severe burns, boiling water or steam?

Solution:

Steam produces severe burns. It is because it is an exothermic reaction that releases a high amount of heat which it had consumed during vaporization.

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