

CHAPTER 1 ACIDS, BASES, BUFFERS

Acids, Bases are defined by four main theories

1. Traditional theory / concept
2. Arrhenius theory
3. Bronsted and Lowry theory
4. Lewis theory

1. Traditional theory / concept

- Acid : are the substance
 - Which converts blue litmus paper to red
 - Having the PH < 7
 - Sour Taste
 - React with bases to form salts and Water
 - Eg :- Hydrochloric acid (HCl)
- Base : are the substance
 - which converts red litmus paper to blue
 - Having the PH >7
 - Bitter taste
 - React with Acids to form salts and water
 - Eg: Sodium Hydroxide (NaOH)

2. Arrhenius theory*

- In 1884 of Svante Arrhenius Also known as, Arrhenius theory of ionization
- Electron dissociation theory
- This theory define acids & bases according to their formation of ions when dissolved in water
- An Acid is a substance that can release hydrogen ion (H⁺) when dissolved in water or A substance which when dissolved in water gives hydrogen ions (H⁺) is known as acid

E.g. Hydrochloric acid (HCl)



Acid - forms H⁺ in water

Chapter 1 study Guide Acids Bases

Robert K. Wismer



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Table of Contents Chapter 1study Guide Acids Bases

1. Understanding the eBook Chapter 1study Guide Acids Bases
 - The Rise of Digital Reading Chapter 1study Guide Acids Bases
 - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 1study Guide Acids Bases
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chapter 1study Guide Acids Bases
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 1study Guide Acids Bases
 - Personalized Recommendations

- Chapter 1study Guide Acids Bases User Reviews and Ratings
- Chapter 1study Guide Acids Bases and Bestseller Lists
- 5. Accessing Chapter 1study Guide Acids Bases Free and Paid eBooks
 - Chapter 1study Guide Acids Bases Public Domain eBooks
 - Chapter 1study Guide Acids Bases eBook Subscription Services
 - Chapter 1study Guide Acids Bases Budget-Friendly Options
- 6. Navigating Chapter 1study Guide Acids Bases eBook Formats
 - ePub, PDF, MOBI, and More
 - Chapter 1study Guide Acids Bases Compatibility with Devices
 - Chapter 1study Guide Acids Bases Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chapter 1study Guide Acids Bases
 - Highlighting and Note-Taking Chapter 1study Guide Acids Bases
 - Interactive Elements Chapter 1study Guide Acids Bases
- 8. Staying Engaged with Chapter 1study Guide Acids Bases
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chapter 1study Guide Acids Bases
- 9. Balancing eBooks and Physical Books Chapter 1study Guide Acids Bases
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chapter 1study Guide Acids Bases
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chapter 1study Guide Acids Bases
 - Setting Reading Goals Chapter 1study Guide Acids Bases
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chapter 1study Guide Acids Bases
 - Fact-Checking eBook Content of Chapter 1study Guide Acids Bases

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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