

Chapter 1: Solutions to Review Problems

Section 1.1

- 1) If the number of moles of gas and the temperature remain constant: $P_1V_1 = P_2V_2$

Substitute: $100 \text{ kPa} \times 2.00 \text{ L} = P_2 \times 3.00 \text{ L} \therefore P_2 = 66.7 \text{ kPa}$

- 2) If the number of moles of gas and the pressure remain constant: $\frac{V_1}{T_1} = \frac{V_2}{T_2}$

Substitute: $\frac{1.00 \text{ L}}{200 \text{ K}} = \frac{V_2}{400 \text{ K}} \therefore V_2 = 2.00 \text{ L}$

- 3) If the number of moles of gas and the volume remain constant: $\frac{P_1}{T_1} = \frac{P_2}{T_2}$

Substitute: $\frac{100 \text{ kPa}}{298 \text{ K}} = \frac{P_2}{398 \text{ K}} \therefore P_2 = 133.6 \text{ kPa}$

- 4) Xe has molar mass 131.3 gmol^{-1} , Ar has molar mass 39.95 gmol^{-1} , Ne has molar mass 20.18 gmol^{-1}

If equal masses of each gas are used, the sample of the gas with the smallest molar mass will contain the greatest number of moles.

Therefore, the sample of Ne will contain the greatest number of moles.

Pressure is directly proportional to the number of moles of gas; so the pressure in the Ne flask is the greatest. (Answer a)

- 5) O_2 has molar mass 32.00 gmol^{-1} , N_2 has molar mass 28.02 gmol^{-1}

If equal masses of each gas are used, the sample of the gas with the smaller molar mass will contain the greater number of moles.

Therefore, the sample of N_2 will contain the greater number of moles, and therefore molecules.

Pressure is directly proportional to the number of moles (or molecules) of gas; so the pressure in the N_2 flask is the greater than the pressure in the O_2 flask. (Answer c)

Chapter 1 solutions Answers

Herbert S. Wilf



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1. Understanding the eBook Chapter 1solutions Answers
 - The Rise of Digital Reading Chapter 1solutions Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 1solutions Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chapter 1solutions Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 1solutions Answers
 - Personalized Recommendations
 - Chapter 1solutions Answers User Reviews and Ratings
 - Chapter 1solutions Answers and Bestseller Lists
5. Accessing Chapter 1solutions Answers Free and Paid eBooks

- Chapter 1 solutions Answers Public Domain eBooks
 - Chapter 1 solutions Answers eBook Subscription Services
 - Chapter 1 solutions Answers Budget-Friendly Options
6. Navigating Chapter 1 solutions Answers eBook Formats
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 - Chapter 1 solutions Answers Compatibility with Devices
 - Chapter 1 solutions Answers Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chapter 1 solutions Answers
 - Highlighting and Note-Taking Chapter 1 solutions Answers
 - Interactive Elements Chapter 1 solutions Answers
 8. Staying Engaged with Chapter 1 solutions Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chapter 1 solutions Answers
 9. Balancing eBooks and Physical Books Chapter 1 solutions Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chapter 1 solutions Answers
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Chapter 1 solutions Answers
 - Setting Reading Goals Chapter 1 solutions Answers
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Chapter 1 solutions Answers
 - Fact-Checking eBook Content of Chapter 1 solutions Answers
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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