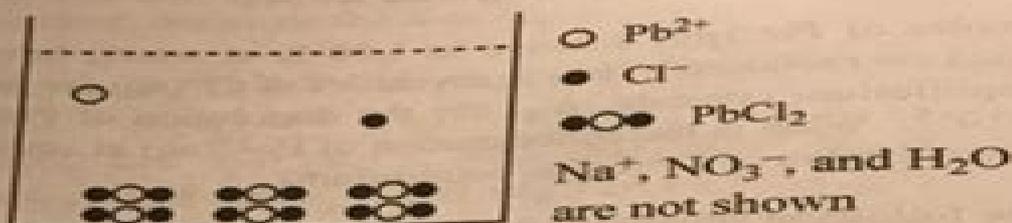


3. Five hundred (500) mL of a $\text{Pb}(\text{NO}_3)_2$ solution and 500 mL of a NaCl solution are added to a large beaker. At equilibrium the result is depicted below:



- a) Which solution prior to mixture, $\text{Pb}(\text{NO}_3)_2$ or NaCl , had the higher concentration?
- b) Explain how you know from the diagram that the solution was not prepared by dissolving PbCl_2 in water.
- c) Given that each symbol, ○ and ●, represents 2.5×10^{-2} moles, calculate the K_{sp} of PbCl_2 at this temperature.

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