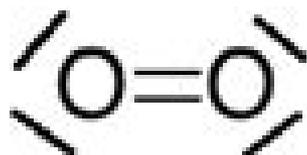


## 4.2 Covalent bonding

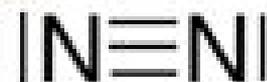
### Multiple covalent bonds

A double covalent bond involves two shared pairs of electrons and a triple covalent bond involves three shared pairs of electrons.



oxygen

Each oxygen atom ( $1s^2 2s^2 2p^4$ ) has six valence electrons. By pairing two of the electrons with two of the electrons from the other oxygen atom to form a **double bond** each atom in the molecule attains the noble gas electron configuration of neon.



nitrogen

Each nitrogen atom ( $1s^2 2s^2 2p^3$ ) has five valence electrons. By pairing three of the electrons with three of the electrons from the other nitrogen atom to form a **triple bond** each atom in the molecule attains the noble gas electron configuration of neon.

# Chem Think Covalent Bonding

**Rachel S Tattersall**



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## **Table of Contents Chem Think Covalent Bonding**

1. Understanding the eBook Chem Think Covalent Bonding
  - The Rise of Digital Reading Chem Think Covalent Bonding
  - Advantages of eBooks Over Traditional Books
2. Identifying Chem Think Covalent Bonding
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an eBook Platform
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chem Think Covalent Bonding
  - Personalized Recommendations
  - eBook User Reviews and Ratings
  - eBook and Bestseller Lists
5. Accessing eBook Free and Paid eBooks
  - eBook Public Domain eBooks
  - eBook Subscription Services
  - eBook Budget-Friendly Options

6. Navigating Chem Think Covalent Bonding eBook Formats
  - ePub, PDF, MOBI, and More
  - Chem Think Covalent Bonding Compatibility with Devices
  - Chem Think Covalent Bonding Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chem Think Covalent Bonding
  - Highlighting and Note-Taking Chem Think Covalent Bonding
  - Interactive Elements Chem Think Covalent Bonding
8. Staying Engaged with Chem Think Covalent Bonding
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chem Think Covalent Bonding
9. Balancing eBooks and Physical Books Chem Think Covalent Bonding
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chem Think Covalent Bonding
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Chem Think Covalent Bonding
  - Setting Reading Goals Chem Think Covalent Bonding
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chem Think Covalent Bonding
  - Fact-Checking eBook Content of Chem Think Covalent Bonding
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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