

Unit 3 ~ Learning Guide Name: Nanya Nita Shetty
Instructions:

Using a pencil, complete the following notes as you work through the related lessons. Show ALL work as is explained in the lessons. You are required to have this package completed BEFORE you write your unit test. Do your best and ask questions if you don't understand anything!

Please print out the Formula Sheet found at the start of the course. You will need this to do assignments and tests!

* Highly Exothermic reaction are less likely to be in equilibrium as they have a greater E_a reverse

Equilibrium:

1. When does equilibrium occur?

Equilibrium occurs when opposing processes occur at equal rates in a closed system at a constant temperature

* Chemical equilibrium is also known as dynamic equilibrium

2. Give an example of two physical process that oppose each other during equilibrium?

A water solution in a closed container can have a physical equilibrium, and the opposing processes would be evaporation and condensation.

3. Complete the ICE chart for this reaction given that 4.0 moles of N₂ gas and 8.0 moles of H₂ gas are placed into a 1.00 L sealed flask and allowed to come to equilibrium. At equilibrium, 2.4 moles of NH₃ are present.

| | N _{2(g)} | + | 3H _{2(g)} | ⇌ | 2 NH _{3(g)} |
|----------------------------------|-------------------|---|--------------------|---|----------------------|
| Initial concentration [I] | 4.0 M | | 8.0 M | | 0 M |
| Change in concentration [C] | -1.2 M | | -3.6 M | | +2.4 M |
| Equilibrium concentration [E] | 2.8 M | | 4.4 M | | 2.4 M |

* Ratio of the balanced equation (1:3:2)

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Organic Chemistry, 12e Binder Ready Version Study Guide & Student Solutions Manual T. W. Graham Solomons, Craig B. Fryhle, Scott A. Snyder, 2016-04-11 This is the Student Study Guide Solutions Manual to accompany Organic Chemistry 12th Edition The 12th edition of Organic Chemistry continues Solomons Fryhle Snyder s tradition of excellence in teaching and preparing students for success in the organic classroom and beyond A central theme of the authors approach to organic chemistry is to emphasize the relationship between structure and reactivity To accomplish this the content is organized in a way that combines the most useful features of a functional group approach with one largely based on reaction mechanisms The authors philosophy is to emphasize mechanisms and their common aspects as often as possible and at the same time use the unifying features of functional groups as the basis for most chapters The structural aspects of the authors approach show students what organic chemistry is Mechanistic aspects of their approach show students how it works And wherever an opportunity arises the authors show students what it does in living systems and the physical world around us Mtel Chemistry (12) Study Guide: Test Prep and Study Questions Trivium Test Prep, 2015-08-17 Think all MTEL Chemistry study guides are the same Think again With easy to understand lessons and practice test questions esigned to maximize your score you ll be ready You don t want to waste time and money retaking an exam You want to accelerate your education not miss opportunities for starting your future career Every year thousands of people think that they are ready for the MTEL Chemistry but realize too late when they get their score back that they were not ready at all They weren t incapable and they certainly did their best but they simply weren t studying the right way There are a variety of methods to prepare for the MTEL Chemistry exam and they get a variety of results Trivium Test Prep s MTEL Chemistry study guide provides the information secrets and confidence needed to get you the score you need the first time around Losing points on the MTEL Chemistry exam can cost you precious time money and effort that you shouldn t have to spend What is in the book In our MTEL Chemistry study guide you get the most comprehensive review of all tested concepts The subjects are easy to understand and have fully explained example questions to ensure that you master the material Best of all we show you how this information will be applied on the real exam MTEL Chemistry practice questions are included so that you can know without a doubt that you are prepared Our study guide is streamlined and concept driven so you get better results through more effective study time Why spend days or even weeks reading through meaningless junk trying to sort out the helpful information from the fluff We give you everything you need to know in a concise comprehensive and effective package Chemistry Study Guide Year 12 Edwin Mark Sayes, 2000-01-01 Study Guide for Chemical Principles [by] Steven S. Zumdahl Steven S. Zumdahl, Paul B. Kelter, 1995 The Study Guide reflects the unique problem solving approach taken by the Chemical Principles text The new edition of the Study Guide includes many new worked out examples

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