

4

Atomic Structure
Reviewsheet

A. Completion

Use this completion exercise to check your understanding of the concepts and terms that are introduced in this chapter. Each blank can be completed with a term, short phrase, or number.

Atoms of each element are 1 from the atoms of all other elements. Dalton theorized that atoms are indivisible, but the discovery of 2 particles changed this theory. We now know that atoms are made up of electrons, which have a 3 charge; 4, which have a positive charge; and 5, which are neutral. The latter two particles are found in the 6 of the atom.

It was 7 who discovered the nucleus of the atom. The nucleus has a 8 charge and it occupies a very small volume of the atom. In contrast, the negatively charged 9 occupy most of the volume of the atom.

The number of 10 in the nucleus of the atom is the atomic 11 of that element. Because atoms are electrically neutral, the number of protons and 12 in an atom are equal. The sum of the 13 and neutrons is the mass number. Atoms of the same element are identical in most respects, but they can differ in the number of 14 in the nucleus. Atoms that have the same number of protons but different mass numbers are called 15.

The 16 of an element is the weighted average of the masses of the isotopes of that element. Two isotopes of sulfur are $^{32}_{16}\text{S}$ and $^{34}_{16}\text{S}$. An atom of the sulfur-32 isotope contains 17 protons and 18 neutrons. The sulfur-34 isotope has 19 protons and 20 neutrons.

1. different 41
2. sub atomic 42
3. negative 42
4. protons 42
5. neutrons 42
6. nucleus 43
7. Rutherford 43
8. positive 43
9. electrons 43
10. protons (elec) 44
11. number 44
12. electrons 44
13. protons (elec) 45
14. neutrons 45
15. isotopes 45
16. atomic mass 47
17. 16 45
18. 16 46
19. 16 46
20. 18 45

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Benjamin Lillard



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Table of Contents Chemistry 8 2 Section Review

1. Understanding the eBook Chemistry 8 2 Section Review
 - The Rise of Digital Reading Chemistry 8 2 Section Review
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry 8 2 Section Review
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry 8 2 Section Review
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry 8 2 Section Review
 - Personalized Recommendations
 - Chemistry 8 2 Section Review User Reviews and Ratings
 - Chemistry 8 2 Section Review and Bestseller Lists
5. Accessing Chemistry 8 2 Section Review Free and Paid eBooks
 - Chemistry 8 2 Section Review Public Domain eBooks
 - Chemistry 8 2 Section Review eBook Subscription Services

- Chemistry 8 2 Section Review Budget-Friendly Options
- 6. Navigating Chemistry 8 2 Section Review eBook Formats
 - ePub, PDF, MOBI, and More
 - Chemistry 8 2 Section Review Compatibility with Devices
 - Chemistry 8 2 Section Review Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry 8 2 Section Review
 - Highlighting and Note-Taking Chemistry 8 2 Section Review
 - Interactive Elements Chemistry 8 2 Section Review
- 8. Staying Engaged with Chemistry 8 2 Section Review
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry 8 2 Section Review
- 9. Balancing eBooks and Physical Books Chemistry 8 2 Section Review
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry 8 2 Section Review
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chemistry 8 2 Section Review
 - Setting Reading Goals Chemistry 8 2 Section Review
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemistry 8 2 Section Review
 - Fact-Checking eBook Content of Chemistry 8 2 Section Review
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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