

Name Key

1. Hydrogen,  $H_2$ , is used as rocket fuel. The hydrogen is burned in oxygen,  $O_2$ , according to the following equation. What is the enthalpy change if 200.0 g of hydrogen,  $H_2$  is burned?



$$\frac{200.0 \text{ g } H_2}{2.016 \text{ g } H_2} \times \frac{1 \text{ mol } H_2}{1} = 99.2 \text{ mol } H_2 \quad \frac{99.2 \text{ mol } H_2}{2 \text{ mol } H_2} \times \frac{-484 \text{ kJ}}{1} = -24,006 \text{ kJ}$$

$-24,006 \text{ kJ}$   
~~24,006 kJ~~

2. Carbon disulfide burns in air, producing carbon dioxide and sulfur dioxide.



What is  $\Delta H$  for the following reaction?



3. Following is the equation for the combustion of acetone:



How much heat will be given off (in kcal) if 4.5 mol of acetone is burned completely?

$$\frac{-853.6 \text{ kcal}}{2 \text{ mol } C_3H_6O} \times 4.5 \text{ mol } C_3H_6O = -1920.6 \text{ kcal}$$

Exothermic

4. The specific heat of iron is  $0.110 \text{ cal} / \text{g} \cdot ^\circ\text{C}$ . How much heat must be added to 5.2 g of iron at  $25^\circ\text{C}$  to raise its temperature to  $225^\circ\text{C}$ ?

$$\text{heat} = q = m c \Delta T = (5.2 \text{ g}) \left( \frac{0.110 \text{ cal}}{\text{g} \cdot ^\circ\text{C}} \right) (225 - 25^\circ\text{C}) = 114.4 \text{ cal}$$

Endothermic

5. Hydrazine,  $N_2H_4$ , is a colorless liquid used as a rocket fuel. What is the enthalpy change for the process in which hydrazine is formed from its elements (reaction (1) below)? Use Hess's Law to calculate  $\Delta H$  for reaction (1) given the data below.



Reverse  
 multiply quantity

# Chemistry 19 3 Practice Problems Answer Key

**M Mark**



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*Calculating Theoretical Yield Is Not Impossible!* Michael S. Leonard, 2025-02-25 Calculating theoretical and percent yield is a fundamental skill for the laboratory This book primarily targets Organic Chemistry Laboratory courses at the high school or college and university level as a supplemental resource to help students master this skill It begins with simple examples from everyday life demonstrates the importance of balancing the equation addresses the role of the mole in these computations discusses different types of liquids considers the role of significant figures and culminates with the planning of syntheses There are suggestions for further reading as well as practice problems and questions to ensure mastery Begins with examples from everyday life that enable students to understand the concepts of theoretical and percent yield before applying those concepts to the laboratory Addresses the necessity of balancing the reaction equation the centrality of the mole in these calculations and the role of significant figures in reporting the answer Explains how to approach the calculations when using neat liquids or solutions The culmination of this text is the use of the same thought processes to plan the amounts of reactants needed for syntheses of desired quantities of product All of the problems in the book include detailed solutions with accompanying text to explain the answers and ancillaries also include suggestions for further reading

Solid-State Chemistry Ashok Kumar Jha, 2023-08-04 Solid state reactions have been growing in importance in the preparation of solids crystal growth and single crystals of elements In this new volume the author discusses in depth recent advances in solid state chemistry the latest techniques of characterization of solids and several new dimensions of solid state problems Solid state reactions their kinetics phase transitions and magnetic properties are discussed in detail along with electrical properties semiconductors metals and insulators with examples and explanations Dielectric piezoelectric and ferroelectric properties of solids have been considered as well Different theories of catalysis and adsorption isotherms are covered and the author includes several experimental findings as well And of course preparation of new solids by taking advantage of recent experimental techniques have also been included in this volume This volume provides a complete description of solid state reactions electrical conductivity characterization of solids organic solids and intercalation between layers of solids The band theories of metals semiconductors and insulators are covered In characterization of solids TGA DTA DSC FTIR and PXRD techniques have been described Preparation and properties of organic solids are also explained Dielectric properties photovoltaic effects piezo electricity ferroelectricity and antiferroelectricity have been dealt with giving relatable examples A number of excitation energy levels giving rise to luminescence and photoluminescence have been explained under the topic optical processes Diamagnetic paramagnetic and ferromagnetic behaviors of solids are also explained

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