

## Stoichiometry Review

**\*\*YOU MUST SHOW ALL WORK AND YOUR ANSWERS MUST INCLUDE THE PROPER NUMBER OF SIG FIGS AND COMPLETE UNITS IN ORDER TO RECEIVE CREDIT FOR THE PROBLEM.\*\***

BALANCE THE FOLLOWING EQUATIONS TO USE IN QUESTIONS 5 – 14:



PERFORM THE FOLLOWING STOICHIOMETRIC CALCULATIONS:

5. Zinc reacts with hydrochloric acid to produce zinc chloride and hydrogen. How many moles of HCl are required to produce 7.50 moles of ZnCl<sub>2</sub>?

$$\frac{7.50 \text{ mol ZnCl}_2}{1 \text{ mol ZnCl}_2} \times \frac{2 \text{ mol HCl}}{1 \text{ mol ZnCl}_2} = 15.0 \text{ mol HCl}$$

6. Copper metal reacts with silver nitrate to form silver and copper(II) nitrate. How many grams of copper are required to form 250 g of silver?

$$\frac{250 \text{ g Ag}}{107.868 \text{ g Ag}} \times \frac{1 \text{ mol Ag}}{2 \text{ mol Ag}} \times \frac{1 \text{ mol Cu}}{1 \text{ mol Cu}} \times 63.546 \text{ g Cu} = 74 \text{ g Cu}$$

7. When aluminum is burned in excess oxygen, aluminum oxide is produced. How many grams of oxygen are required to produce 0.75 moles of Al<sub>2</sub>O<sub>3</sub>?

$$\frac{0.75 \text{ mol Al}_2\text{O}_3}{2 \text{ mol Al}_2\text{O}_3} \times \frac{3 \text{ mol O}_2}{1 \text{ mol O}_2} \times 31.998 \text{ g O}_2 = 36 \text{ g O}_2$$

8. How many grams of iron(III) chloride are produced when 15.3 g of iron react with excess chlorine gas?

$$\frac{15.3 \text{ g Fe}}{55.845 \text{ g Fe}} \times \frac{1 \text{ mol Fe}}{2 \text{ mol Fe}} \times \frac{2 \text{ mol FeCl}_3}{1 \text{ mol FeCl}_3} \times 162.204 \text{ g FeCl}_3 = 44.4 \text{ g FeCl}_3$$

9. Copper metal reacts with silver nitrate to form silver and copper(II) nitrate. How many moles of silver will be produced from 3.65 moles of silver nitrate?

$$\frac{3.65 \text{ mol AgNO}_3}{2 \text{ mol AgNO}_3} \times \frac{2 \text{ mol Ag}}{1 \text{ mol AgNO}_3} = 3.65 \text{ mol Ag}$$

# Chemical Reactions And Stoichiometry Review Answers

**Mary Millhollon, Richard H. Langley**



## **Chemical Reactions And Stoichiometry Review Answers:**

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*Ebook: Chemistry: The Molecular Nature of Matter and Change* Silberberg, 2015-01-16 Ebook Chemistry The Molecular Nature of Matter and Change

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**Modern Chemistry** Holt Rinehart & Winston, Holt, Rinehart and Winston Staff, 2001

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*A Complete Preparation for the M.C.A.T.* James L. Flowers, 1991

**Student Study Guide to accompany Chemistry** Martin Silberberg, Libby Weberg, 2005-01-06

Columbia Review High-yield General Chemistry Stephen D. Bresnick, 1996 A study guide in question and answer format for basic chemistry

*General, Organic, and Biological Chemistry* Drew H. Wolfe, 1986

*General Chemistry, Study Guide* James E. Brady, 1990-03-01 The Fifth Edition retains the pedagogical strengths that made the previous editions so popular and has been updated reorganized and streamlined Changes include more accessible introductory chapters with greater stress on the logic of the periodic table earlier introduction of redox reactions greater emphasis on the concept of energy a new section on Lewis structures earlier introduction of the ideal gas law and a new development of thermodynamics Each chapter ends with review questions and problems

*Merrill Chemistry* Robert C. Smoot, Smoot, Richard G. Smith, Jack

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