

Chapter 10 Chemical Quantities Review

Part 1: Vocabulary

Label the structure corresponding to the term in the parentheses.

- 1.** mole: gram-molecular weight
- 2.** empirical formula: simplest formula
- 3.** molar mass: relative molecular weight
- 4.** 10^{-24} : unit of mass
- 5.** a description of the atoms present in a substance: composition
- 6.** the lowest whole-number ratio of the atoms of the elements in a compound: empirical formula
- 7.** formula structure: the way in which atoms are joined: structural formula
- 8.** the number of moles of a compound: number of moles
- 9.** the formula representing $n \times 10^{-24}$ of 10^{23} representative particles of a substance: mole
- 10.** 10^{23} particles: Avogadro's number
- 11.** the mass of a compound of a certain mass of an element: relative molecular weight (R.M.W.)

Part 2: Section Concept Review

10.1 The Mole: A Fundamental

Completion Use the completion method to check your knowledge of the terms and facts concerning the concepts involved in the section. Each blank space represents a term that's given in outline.

- Knowing what mole counting means, it's easy to calculate the number of particles in a sample having a certain mass. Avogadro's number represents the number of particles in one mole of a substance. It's Avogadro's number.
- The number of moles of a compound, element, and together the number of moles of the substances in the compound, when you substitute the values for atoms, you obtain the molar mass of the compound. These are Avogadro's number representative particles the mass of one substance.

True or False

- 10.1** A mole of any substance has a mass of 1 gram.
- 10.2** The number of particles in a mole of a substance is the same for all substances.
- 10.3** The number of particles in a mole of a substance is the same for all substances.
- 10.4** The mass of a mole of a substance is the same for all substances.

Problems

Chemical Quantities Vocabulary Review Answer Key

G Thomas



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