

Chapter 10 Chemical Quantities Review

Part 1: Vocabulary

Label the structure corresponding to the term in the parentheses.

- 1.** mole (molar mass) molar mass
- 2.** empirical formula (molecular formula) empirical formula
- 3.** molar mass (molar weight) molar mass
- 4.** molar mass (molar weight) molar mass
- 5.** molar mass (molar weight) molar mass
- 6.** molar mass (molar weight) molar mass
- 7.** molar mass (molar weight) molar mass
- 8.** molar mass (molar weight) molar mass
- 9.** molar mass (molar weight) molar mass
- 10.** molar mass (molar weight) molar mass
- 11.** molar mass (molar weight) molar mass
- 12.** molar mass (molar weight) molar mass
- 13.** molar mass (molar weight) molar mass
- 14.** molar mass (molar weight) molar mass
- 15.** molar mass (molar weight) molar mass
- 16.** molar mass (molar weight) molar mass
- 17.** molar mass (molar weight) molar mass
- 18.** molar mass (molar weight) molar mass
- 19.** molar mass (molar weight) molar mass
- 20.** molar mass (molar weight) molar mass

Part 2: Section Concept Review

10.1 The Mole: A Fundamental

Complete the following exercises to test your knowledge of the terms and your understanding of the concepts involved with a mole of a substance.

- 1.** How many atoms are there in 1 mole of a substance? 6.022 x 10²³
- 2.** How many molecules are there in 1 mole of a substance? 6.022 x 10²³
- 3.** How many atoms are there in 1 mole of a substance? 6.022 x 10²³
- 4.** How many molecules are there in 1 mole of a substance? 6.022 x 10²³
- 5.** How many atoms are there in 1 mole of a substance? 6.022 x 10²³
- 6.** How many molecules are there in 1 mole of a substance? 6.022 x 10²³
- 7.** How many atoms are there in 1 mole of a substance? 6.022 x 10²³
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- 16.** How many molecules are there in 1 mole of a substance? 6.022 x 10²³
- 17.** How many atoms are there in 1 mole of a substance? 6.022 x 10²³
- 18.** How many molecules are there in 1 mole of a substance? 6.022 x 10²³
- 19.** How many atoms are there in 1 mole of a substance? 6.022 x 10²³
- 20.** How many molecules are there in 1 mole of a substance? 6.022 x 10²³

Practice Problems

- 1.** Calculate the molar mass of a substance.
- 2.** Calculate the molar mass of a substance.
- 3.** Calculate the molar mass of a substance.
- 4.** Calculate the molar mass of a substance.
- 5.** Calculate the molar mass of a substance.
- 6.** Calculate the molar mass of a substance.
- 7.** Calculate the molar mass of a substance.
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- 11.** Calculate the molar mass of a substance.
- 12.** Calculate the molar mass of a substance.
- 13.** Calculate the molar mass of a substance.
- 14.** Calculate the molar mass of a substance.
- 15.** Calculate the molar mass of a substance.
- 16.** Calculate the molar mass of a substance.
- 17.** Calculate the molar mass of a substance.
- 18.** Calculate the molar mass of a substance.
- 19.** Calculate the molar mass of a substance.
- 20.** Calculate the molar mass of a substance.

Answers

Chemical Quantities Ch Vocab Review Key

DP Hallahan



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