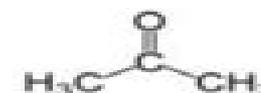


Unit 0 – Chemistry Review

Section 3 – Chemical Formulas



Molecular (Covalent) Compounds

- What information can we tell about a compound given its empirical formula?
We can tell the smallest whole-number ratio of all the atoms in a given compound.
- What additional information is conveyed when we know its molecular formula?
We can tell the mass of the molecule.
- What additional information do we know when we see its structural formula?
We can tell what each atom is bonded to.
- Two compounds have the same empirical formula; one is a gas at room temperature and the other is a liquid. How is this possible?
Just because they have the same empirical formula does not mean they are the same compound. Small changes in atom arrangement and quantity can result in large differences in physical properties.
- Determine the molecular and empirical formula of the following: **Molecular first, then empirical.**
 - The organic solvent benzene that has six carbon atoms and six hydrogen atoms **C₆H₆, CH**
 - The compound silicon tetrachloride, a substance used to make computer chips **SiCl₄, SiCl₄**
 - The reactive substance diborane, a compound with two boron atoms and six hydrogen atoms **B₂H₆, BH₃**
 - The sugar glucose, a substance with six carbon, twelve hydrogen, and six oxygen atoms **C₆H₁₂O₆, CH₂O**
- Write down the molecular and structural formulas represented by the following models:



Notice how some are written slightly different. They are all correct representations even though one is more condensed and one shows the valence electrons.

Ionic Compounds

- Which of the following compounds below likely represents an ionic compound and which represents a covalent compound? Explain your answer.
The first one is likely an ionic compound because of the repeating pattern of positive and negative ions.

- Predict the charge of the most stable ion each element below is capable of forming:

a. Mg **2+**
b. Al **3+**

c. S **2-**
d. K **1+**

e. F **1-**
f. O **2-**



g. N **3-**
h. Na **1+**

- Fill in the table for the compound made from each cation and anion pair. Use parentheses when necessary.

	Cl ⁻	NO ₃ ⁻	S ²⁻	CO ₃ ²⁻	N ³⁻	PO ₄ ³⁻	OH ⁻
Na ⁺	NaCl	NaNO ₃	Na ₂ S	Na ₂ CO ₃	Na ₃ N	Na ₃ PO ₄	NaOH
NH ₄ ⁺	NH ₄ Cl	NH ₄ NO ₃	(NH ₄) ₂ S	(NH ₄) ₂ CO ₃	(NH ₄) ₃ N	(NH ₄) ₃ PO ₄	NH ₄ OH
Sn ²⁺	SnCl ₂	Sn(NO ₃) ₂	SnS	SnCO ₃	Sn ₃ N ₂	Sn ₃ (PO ₄) ₂	Sn(OH) ₂
Hg ₂ ²⁺	Hg ₂ Cl ₂	Hg ₂ (NO ₃) ₂	Hg ₂ S	Hg ₂ CO ₃	(Hg ₂) ₂ N ₂	(Hg ₂) ₂ (PO ₄) ₂	Hg ₂ (OH) ₂
Al ³⁺	AlCl ₃	Al(NO ₃) ₃	Al ₂ S ₃	Al ₂ (CO ₃) ₃	AlN	AlPO ₄	Al(OH) ₃
Sn ⁴⁺	SnCl ₄	Sn(NO ₃) ₄	SnS ₂	Sn(CO ₃) ₂	Sn ₃ N ₄	Sn ₃ (PO ₄) ₄	Sn(OH) ₄

NOTE: ONE of the above formulas is incorrect. If you can spot it, send me an e-mail. The first one to find it and email me the answer will receive extra credit!

Chemical Formulas And Compounds Chapter Review

Section 1

O García



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