

Previous Years AIPMT/NEET & AIIMS Questions

1. For reaction $2\text{NOCl}(g) \rightleftharpoons 2\text{NO}(g) + \text{Cl}_2(g)$, K_c at 427°C is $3 \times 10^{-5} \text{ L}^{-1} \text{ mol}$. The value of K_p is nearly :- [AIIMS-2005]

- (1) 7.50×10^{-5} (2) 2.50×10^{-2} (3) 2.50×10^{-4} (4) 1.75×10^{-4}

2. Equilibrium constants K_1 and K_2 for the following equilibrium : [AIPMT-2005]

$\text{NO}(g) + \frac{1}{2} \text{O}_2(g) \xrightleftharpoons{K_1} \text{NO}_2(g)$ and $2\text{NO}_2(g) \xrightleftharpoons{K_2} 2\text{NO}(g) + \text{O}_2(g)$ are related as :-

- (1) $K_2 = \frac{1}{K_1}$ (2) $K_2 = \frac{K_1}{2}$ (3) $K_2 = \frac{1}{K_1^2}$ (4) $K_2 = K_1^2$

3. For the reaction :



$\Delta H = -170.8 \text{ kJ mol}^{-1}$

Which of the following statements is not true :- [AIPMT-2006]

- (1) At equilibrium, the concentrations of $\text{CO}_2(g)$ and $\text{H}_2\text{O}(l)$ are not equal
 (2) The equilibrium constant for the reaction is given by $K_p = \frac{[\text{CO}_2]}{[\text{CH}_4][\text{O}_2]}$
 (3) Addition of $\text{CH}_4(g)$ or $\text{O}_2(g)$ at equilibrium will cause a shift to the right
 (4) The reaction is exothermic

4. The following equilibrium are given



The equilibrium constant of the reaction

$2\text{NH}_3 + \frac{5}{2} \text{O}_2 \rightleftharpoons 2\text{NO} + 3 \text{H}_2\text{O}$, in terms of K_1 , K_2 and K_3 is : [AIPMT-2007]

- (1) $\frac{K_1 K_2}{K_3}$ (2) $\frac{K_1 K_2^2}{K_3}$ (3) $\frac{K_2 K_3^2}{K_1}$ (4) $K_1 K_2 K_3$

5. The dissociation equilibrium of a gas AB_2 can be represented as : $2\text{AB}_2(g) \rightleftharpoons 2\text{AB}(g) + \text{B}_2(g)$ The degree of dissociation is 'x' and is small compared to 1. The expression relating the degree of dissociation (x) with equilibrium constant K_p and total pressure P is :

- (1) $(2K_p/P)^{1/2}$ (2) $(2K_p/P)^{1/2}$ [AIPMT-2008]
 (3) (K_p/P) (4) $(2K_p/P)$

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