

6**Chemical Quantities****Downloaded****A. Connections**

This diagnostic exercise is meant to check your knowledge of the terms and your understanding of the concepts introduced in this chapter. Each blank area has multiple-choice options, three phrases, or numbers.

Chemists define units of measuring, of mass, and of volume as a single quantity called the 1. The unit measures the "amount of substance." One way to measure the amount of a substance is to count the number of 2 of that substance. The number of representative particles in a mole of a substance is equal 3. This number is 4. For example, a mole of nitrogen contains 6.02×10^{23} 5 of nitrogen.

Chemists define gram atomic mass as the weight of grams of an element whose mass is equal to the 6 atomic mass. This is equal to 11.0 g of carbon, equal to the mass number of carbon as 7 g of oxygen atoms.

To find the number of moles of a substance represented, we will require the 8 of the substance and its molecular mass. When we substitute the molecular mass, we obtain the 9 of the compound. If we divide molecular mass by its representative number, then we get the number of particles in the compound. The number of particles in a mole is called as 10. The term gram formula mass and 11 are the same as well as the mass of a mole of any substance or compound.

The gram formula mass is defined as the sum of masses of a substance as made up of atoms of a substance together. It depends on the number of moles as a multiple of its representative number. For example, the mass of 2 moles of a compound is 12. The number of moles of a substance is obtained by the ratio 13 for the substance.

1 mass2 representative particles3 Avogadro's Number4 6.02×10^{23} 5 molecules6 atomic mass7 128 gram atomic masses9 gram molecular mass10 gram formula mass11 Molar mass12 conversion factor13 chemical formula

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Geoffrey Arthur Harrison



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