

Chapter 3

Chemical Foundations: Elements, Atoms, and Ions

I. Matching

Match the description in Column B with the correct term in Column A. Write the letter in the blank provided. Each term matches with only one description, so be sure to choose the best description for each term.

Column A

- _____ 1. electron
- _____ 2. nucleus
- _____ 3. isotopes
- _____ 4. atomic number
- _____ 5. mass number
- _____ 6. ion
- _____ 7. neutron
- _____ 8. cation
- _____ 9. proton
- _____ 10. anion

Column B

- A. sum of the number of protons and the number of neutrons in the nucleus of an atom
- B. positively charged subatomic particle
- C. dense, positively charged center of atom
- D. subatomic particle with zero charge
- E. atoms with same number of protons but different numbers of neutrons
- F. negatively charged ion
- G. number of protons in the nucleus of an atom
- H. atom with a charge
- I. positively charged ion
- J. negatively charged subatomic particle

II. Multiple Choice

Choose the one best answer and write its letter in the blank.

_____ 11. Which of the following is a compound?

a) Fe

b) CO

c) O₂

d) K

Chemical Foundations Answers

RJ Shavelson



Chemical Foundations Answers:

Aims and Purposes of the Chemical Foundation, Incorporated Alexander Mitchell Palmer, Francis Patrick Garvan, 1919

Drug and Chemical Markets, 1918 **Drug & Chemical Markets**, 1926 **Chemical, Color and Oil Record**, 1922

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Electrochemical and Metallurgical Industry, 1922
Science, 1926 Chemist and Druggist, 1922 **American Dyestuff Reporter**, 1924 Includes proceedings of American Association of Textile Chemists and Colorists **Chemical Engineer**, 1919 **Chemical Age**, 1919
Proceedings of the Semi-annual [or, Annual] Meeting of N.E. Cotton Manufacturers' Association National Association of Cotton Manufacturers, 1919 **Posselt's Textile Journal**, 1920 **Journal of Pharmaceutical Sciences**, 1922 Vols for 1912-45 include proceedings of the association's annual meeting Journal of the American Pharmaceutical Association American Pharmaceutical Association, 1922 **Textile World Journal**, 1919 *Proceedings of the High School Conference of ...* Horace Adelbert Hollister, 1919 Study Guide for Introductory Chemistry, a Foundation, Introductory Chemistry, Basic Chemistry Iris Stovall, Steven S. Zumdahl, 1996

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web oct 31 2023 abstractpurpose fibroblast activation protein fap is a promising target for tumor treatment in this study we aimed to investigate the safety and efficacy of the albumin binder conjugated fap targeted radiopharmaceutical ^{177}Lu eb fapi ^{177}Lu lnc1004 in patients with metastatic radioiodine refractory thyroid cancer mrair tc experimental

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web 1 consider the reaction $\text{I}_2\text{O}_5 + 5\text{CO} \rightarrow 5\text{CO}_2 + \text{I}_2$ g a 80.0 grams of iodine v oxide I_2O_5 reacts with 28.0 grams of carbon monoxide co determine the mass of iodine I_2 which could be produced b if in the above situation only 0.160 moles of iodine I_2

was produced i what mass of iodine was produced

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web limiting reactant reactant present in an amount lower than required by the reaction stoichiometry thus limiting the amount of product generated percent yield measure of the efficiency of a reaction expressed as a percentage of the theoretical yield

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$$\text{Cu} + 2\text{AgNO}_3 \rightarrow \text{Cu}(\text{NO}_3)_2 + 2\text{Ag}$$

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web start with 8 3 mole h2s and multiply it with 2 mole h2o over 2 mole h2s from the equation 8 3 mole h2s x 2 mole h2o / 2 mole h2s x 18g h2o = 149 4g h2o is your theoretical yield so percent yield = 137 1g h2o actual yield / 149 4g h2o x 100 = 91 77 problems 6 8 are similar to problem 5

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web 1 csun edu hcchm001 introchemhandouts html limiting reagents theoretical actual and percent yields example of a limiting reagent problem how many grams of nh 3 can be be produced theoretically from the reaction of 5 0 g of n what is the limiting reagent if 8 52 g are actually formed what is the

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web percent yield is a measure of the efficiency of the experimental design yield efficiency mass of product obtained
calculated mass of the product expected $\times 100$ in this reaction a yield 80 is anticipated

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web a write a balanced equation for the reaction $\text{Al}_2\text{O}_3 + 4\text{Al} + 3\text{O}_2 \rightarrow 2\text{Al}_2\text{O}_3$ b determine the theoretical yield of Al_2O_3 160 g
 $\text{Al} \times 1 \text{ mol Al} \times 2 \text{ mol Al}_2\text{O}_3 \times 101.964 \text{ g Al}_2\text{O}_3 = 302.3 \text{ g Al}_2\text{O}_3$ c determine the percent yield
yield actual $\times 100 = \frac{260.0 \text{ g}}{302.3 \text{ g}} \times 100 = 86.01\%$ theoretical 302.3 g

7 2 limiting reagent and reaction yields chem 1114 - Jul 02 2022

web 7 na 2 c 2 o 4 is the limiting reactant percent yield 86.6% only four molecules can be made 9 this amount cannot be
weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12
c is the limiting reagent 4.33 g of h₂ are left over 13

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from a set of reactants calculate how much product will be produced from the limiting reagent calculate how much reactant s
remains when the reaction is complete

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web the percent yield is calculated as follows $\text{percent yield} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100$
yield theoretical $\times \frac{\text{actual yield}}{\text{theoretical yield}} = 100$ based on this definition we would expect a percent yield

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web may 20 2018 learning objectives identify the limiting reactant limiting reagent in a given chemical reaction calculate
how much product will be produced from the limiting reactant calculate how much reactant s remains when the reaction is
complete

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 CH_3OH can be produced by the following reaction $2\text{H}_2 + \text{CO} \rightarrow \text{CH}_3\text{OH}$ calculate the theoretical yield of CH_3OH if 68.5 g of co is
reacted with 8.6 g of h₂

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problems from a balanced chemical equation 3 sto 5 differentiate between the actual yield and theoretical yield of a chemical
reaction

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web jun 2 2020 the percent yield of a reaction is the ratio of the actual yield to the theoretical yield multiplied by 100 to give a percentage text percent yield text actual yield g over text theoretical yield g times 100 label 3 7 3