

## Section 2: Practicing equation balancing

Before you can write a balanced equation for a problem which asks you to predict the products of a reaction, you need to know how to balance an equation. Because some of you may not fully remember how to balance an equation, here are some practice problems:

- 1)  $\_\_ \text{C}_6\text{H}_6 + \_\_ \text{O}_2 \rightarrow \_\_ \text{H}_2\text{O} + \_\_ \text{CO}_2$
- 2)  $\_\_ \text{NaI} + \_\_ \text{Pb}(\text{SO}_4)_2 \rightarrow \_\_ \text{PbI}_4 + \_\_ \text{Na}_2\text{SO}_4$
- 3)  $\_\_ \text{NH}_3 + \_\_ \text{O}_2 \rightarrow \_\_ \text{NO} + \_\_ \text{H}_2\text{O}$
- 4)  $\_\_ \text{Fe}(\text{OH})_3 \rightarrow \_\_ \text{Fe}_2\text{O}_3 + \_\_ \text{H}_2\text{O}$
- 5)  $\_\_ \text{HNO}_3 + \_\_ \text{Mg}(\text{OH})_2 \rightarrow \_\_ \text{H}_2\text{O} + \_\_ \text{Mg}(\text{NO}_3)_2$
- 6)  $\_\_ \text{H}_3\text{PO}_4 + \_\_ \text{NaBr} \rightarrow \_\_ \text{HBr} + \_\_ \text{Na}_3\text{PO}_4$
- 7)  $\_\_ \text{C} + \_\_ \text{H}_2 \rightarrow \_\_ \text{C}_3\text{H}_8$
- 8)  $\_\_ \text{CaO} + \_\_ \text{MnI}_4 \rightarrow \_\_ \text{MnO}_2 + \_\_ \text{CaI}_2$
- 9)  $\_\_ \text{Fe}_2\text{O}_3 + \_\_ \text{H}_2\text{O} \rightarrow \_\_ \text{Fe}(\text{OH})_3$
- 10)  $\_\_ \text{C}_2\text{H}_2 + \_\_ \text{H}_2 \rightarrow \_\_ \text{C}_2\text{H}_6$
- 11)  $\_\_ \text{VF}_5 + \_\_ \text{HI} \rightarrow \_\_ \text{V}_2\text{I}_{10} + \_\_ \text{HF}$
- 12)  $\_\_ \text{OsO}_4 + \_\_ \text{PtCl}_4 \rightarrow \_\_ \text{PtO}_2 + \_\_ \text{OsCl}_6$
- 13)  $\_\_ \text{CF}_4 + \_\_ \text{Br}_2 \rightarrow \_\_ \text{CBr}_4 + \_\_ \text{F}_2$
- 14)  $\_\_ \text{Hg}_2\text{I}_2 + \_\_ \text{O}_2 \rightarrow \_\_ \text{Hg}_2\text{O} + \_\_ \text{I}_2$
- 15)  $\_\_ \text{Y}(\text{NO}_3)_3 + \_\_ \text{GaPO}_4 \rightarrow \_\_ \text{YPO}_4 + \_\_ \text{Ga}(\text{NO}_3)_3$

# Chemfiesta Section 2 Practicing Equation Balancing Answers

**Kraaya Publishing**



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