

Name: \_\_\_\_\_



## Balancing Chemical Equations

Directions: Balance each of the following chemical equations.

- 1)  $\_\_\_ \text{C}_{30}\text{H}_{72} + \_\_\_ \text{O}_2 \rightarrow \_\_\_ \text{CO}_2 + \_\_\_ \text{H}_2\text{O}$
- 2)  $\_\_\_ \text{Al}(\text{OH})_3 + \_\_\_ \text{HBr} \rightarrow \_\_\_ \text{AlBr}_3 + \_\_\_ \text{H}_2\text{O}$
- 3)  $\_\_\_ \text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_3 + \_\_\_ \text{O}_2 \rightarrow \_\_\_ \text{CO}_2 + \_\_\_ \text{H}_2\text{O}$
- 4)  $\_\_\_ \text{C}_7\text{H}_8 + \_\_\_ \text{O}_2 \rightarrow \_\_\_ \text{CO}_2 + \_\_\_ \text{H}_2\text{O}$
- 5)  $\_\_\_ \text{Li} + \_\_\_ \text{AlCl}_3 \rightarrow \_\_\_ \text{LiCl} + \_\_\_ \text{Al}$
- 6)  $\_\_\_ \text{C}_3\text{H}_6 + \_\_\_ \text{O}_2 \rightarrow \_\_\_ \text{CO}_2 + \_\_\_ \text{H}_2\text{O}$
- 7)  $\_\_\_ \text{NH}_4\text{OH} + \_\_\_ \text{H}_3\text{PO}_4 \rightarrow \_\_\_ (\text{NH}_4)_3\text{PO}_4 + \_\_\_ \text{H}_2\text{O}$
- 8)  $\_\_\_ \text{Rb} + \_\_\_ \text{P} \rightarrow \_\_\_ \text{Rb}_3\text{P}$
- 9)  $\_\_\_ \text{CH}_4 + \_\_\_ \text{O}_2 \rightarrow \_\_\_ \text{CO}_2 + \_\_\_ \text{H}_2\text{O}$
- 10)  $\_\_\_ \text{Al}(\text{OH})_3 + \_\_\_ \text{H}_2\text{SO}_4 \rightarrow \_\_\_ \text{Al}_2(\text{SO}_4)_3 + \_\_\_ \text{H}_2\text{O}$
- 11)  $\_\_\_ \text{Na} + \_\_\_ \text{Cl}_2 \rightarrow \_\_\_ \text{NaCl}$
- 12)  $\_\_\_ \text{Rb} + \_\_\_ \text{S}_8 \rightarrow \_\_\_ \text{Rb}_2\text{S}$
- 13)  $\_\_\_ \text{H}_3\text{PO}_4 + \_\_\_ \text{Ca}(\text{OH})_2 \rightarrow \_\_\_ \text{Ca}_3(\text{PO}_4)_2 + \_\_\_ \text{H}_2\text{O}$
- 14)  $\_\_\_ \text{NH}_3 + \_\_\_ \text{HCl} \rightarrow \_\_\_ \text{NH}_4\text{Cl}$
- 15)  $\_\_\_ \text{Li} + \_\_\_ \text{H}_2\text{O} \rightarrow \_\_\_ \text{LiOH} + \_\_\_ \text{H}_2$
- 16)  $\_\_\_ \text{Ca}_3(\text{PO}_4)_2 + \_\_\_ \text{SiO}_2 + \_\_\_ \text{C} \rightarrow \_\_\_ \text{CaSiO}_3 + \_\_\_ \text{CO} + \_\_\_ \text{P}$
- 17)  $\_\_\_ \text{NH}_3 + \_\_\_ \text{O}_2 \rightarrow \_\_\_ \text{N}_2 + \_\_\_ \text{H}_2\text{O}$
- 18)  $\_\_\_ \text{FeS}_2 + \_\_\_ \text{O}_2 \rightarrow \_\_\_ \text{Fe}_2\text{O}_3 + \_\_\_ \text{SO}_2$

# Chemfiesta Lab Answers

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