

Ions and Ionic Compounds

(How are ionic compounds formed?)

Model 1: Common charges (by group) on elements when in ionic compounds.

Charge		← variable charge →										⊗				⊗	
Grp #																	
1																	2
H																	He
3	4											5	6	7	8	9	10
Li	Be											B	C	N	O	F	Ne
11	12											13	14	15	16	17	18
Na	Mg											Al	Si	P	S	Cl	Ar
19	20	21	22	23	24	25	26	27	28	29	30	31	32	33	34	35	36
K	Ca	Sc	Ti	V	Cr	Mn	Fe	Co	Ni	Cu	Zn	Ga	Ge	As	Se	Br	Kr
37	38	39	40	41	42	43	44	45	46	47	48	49	50	51	52	53	54
Rb	Sr	Y	Zr	Nb	Mo	Tc	Ru	Rh	Pd	Ag	Cd	In	Sn	Sb	Te	I	Xe
55	56	57	72	73	74	75	76	77	78	79	80	81	82	83	84	85	86
Cs	Ba	La	Hf	Ta	W	Re	Os	Ir	Pt	Au	Hg	Tl	Pb	Bi	Po	At	Rn
87	88	89	104	105	106	107	108	109	110	111	112		114		116		
Fr	Ra	Ac	Rf	Db	Sg	Bh	Hs	Mt	Ds	Rg	Cn		Fl		Lv		

A cation has a positive charge.

An anion has a negative charge.

Critical Thinking Questions:

- Recall the shell model of the atom.
 - How many electrons fill the first shell? _____
 - How many electrons fill the second shell? _____
- How many valence electrons are found in atoms in
 - group 1? _____
 - group 13? _____
 - group 16? _____
- When atoms form ionic compounds, they often fill their outer valence shell (**octet rule**) by either gaining or losing 1, 2, or 3 electrons. Based on your answer to CTQ 2, indicate how many electrons these atoms would either gain or lose to achieve an octet.
 - group 1? _____
 - group 13? _____
 - group 16? _____
- What would be the resulting charge on the ions created in CTQ 3 when electrons are gained or lost?
 - group 1? _____
 - group 13? _____
 - group 16? _____
- Based on your answer to CTQ 4, do metals typically form **anions** or **cations** (circle one)?
- Work with your team and write charges in the empty boxes above each group (skip the transition metals, groups 3-12) at the top of the periodic table in Model 1.
- Do nonmetals typically form **anions** or **cations** (circle one)?

Chemactivity 11 Answers

Andrei Straumanis



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Table of Contents Chemactivity 11 Answers

1. Understanding the eBook Chemactivity 11 Answers
 - The Rise of Digital Reading Chemactivity 11 Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemactivity 11 Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemactivity 11 Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemactivity 11 Answers
 - Personalized Recommendations
 - Chemactivity 11 Answers User Reviews and Ratings
 - Chemactivity 11 Answers and Bestseller Lists
5. Accessing Chemactivity 11 Answers Free and Paid eBooks

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 - Chemactivity 11 Answers Budget-Friendly Options
6. Navigating Chemactivity 11 Answers eBook Formats
- ePub, PDF, MOBI, and More
 - Chemactivity 11 Answers Compatibility with Devices
 - Chemactivity 11 Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Chemactivity 11 Answers
 - Highlighting and Note-Taking Chemactivity 11 Answers
 - Interactive Elements Chemactivity 11 Answers
8. Staying Engaged with Chemactivity 11 Answers
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemactivity 11 Answers
9. Balancing eBooks and Physical Books Chemactivity 11 Answers
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemactivity 11 Answers
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chemactivity 11 Answers
- Setting Reading Goals Chemactivity 11 Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemactivity 11 Answers
- Fact-Checking eBook Content of Chemactivity 11 Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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