

SACRED HEART CENTRAL SCHOOL (SENIOR SECONDARY)

VILLUPURAM

CHEMISTRY - 043

STD: XII

CHAPTERS: 3,8,10,12&14

TIME: 3.00 HRS

HALF PORTION TEST (03/01/23)

MAXIMUM MARKS: 70

SECTION A

The following questions are multiple-choice questions with one correct answer. Each question carries 1 mark. There is no internal choice in this section

- An electrochemical cell behave like an electrolytic cell when
(a) $E_{\text{cell}} = E_{\text{external}}$ (b) $E_{\text{cell}} = 0$ (c) $E_{\text{external}} > E_{\text{cell}}$ (d) $E_{\text{external}} < E_{\text{cell}}$
- Which of the following statement about transition element is not correct?
(a) They show variable oxidation states. (b) They exhibit diamagnetic and paramagnetic properties.
(c) All ions are coloured. (d) They exhibit catalytic property.
- Which of the following is most preferred to prepare alkylchloride from alcohol?
(a) Concentrated HCl (b) SOCl_2 (c) PCl_5 (d) PCl_3
- Carbonyl group undergoes
(a) Electrophilic addition reaction. (b) Nucleophilic addition reactions (c) Both (d) None
- Which of the following reaction confirms the presence of carbonyl group ($> \text{C}=\text{O}$) in glucose?
(a) Reaction with HI (b) Reaction with hydroxylamine (c) Reaction with HCN (d) Both (b) and (c)
- How many coulombs are required for the oxidation of 1 mol of H_2O to O_2 ?
(a) 391000 C (b) 193000 C (c) 319000 C (d) 390100 C
- Which one of following does not correctly represent the correct order of the property indicated against it ?
(a) $\text{Ti} < \text{V} < \text{Cr} < \text{Mn}$, increasing number of oxidation states
(b) $\text{Ti} < \text{V} < \text{Mn} < \text{Cr}$, increasing second ionisation enthalpy
(c) $\text{Ti} < \text{V} < \text{Cr} < \text{Mn}$, increasing melting point
(d) $\text{Ti}^{3+} < \text{V}^{3+} < \text{Cr}^{3+} < \text{Mn}^{3+}$, increasing magnetic moment
- The reagents for the following conversion is/are $\text{Br}(\text{CH}_2)_2\text{Br} \rightarrow \text{C}_2\text{H}_2$
(a) $\text{Zn}/\text{CH}_3\text{OH}$ (b) alc. KOH followed by NaNH_2 (c) aq. KOH followed by NaNH_2 (d) alc. KOH
- Fructose reduces Tollens' reagent due to
(a) Primary alcoholic group
(b) Asymmetric carbons
(c) Secondary alcoholic groups
(d) Enolisation of fructose followed by conversion to aldehyde by base
- Which of the following naturally occurring α - aminoacids is optically inactive?
(a) Glycine (b) Alanine (c) Leucine (d) Valine
- Alkyl halides undergoing nucleophilic bimolecular substitution involve
(a) retention of configuration (b) racemic mixture (c) invert configuration (d) formation of carbocation
- The electrode potential data are given below
 $\text{Fe}^{3+} + e \rightarrow \text{Fe}^{2+}$ $E^\circ = +0.77\text{V}$; $\text{Al}^{3+} + 3e \rightarrow \text{Al}$ $E^\circ = -1.66\text{V}$; $\text{Br}_2 + 2e \rightarrow 2\text{Br}^-$ $E^\circ = +1.80\text{V}$
Based on the data, the reducing power of Fe^{2+} , Al, Br^- will increase in the order.
(a) $\text{Br}^- < \text{Fe}^{2+} < \text{Al}$ (b) $\text{Fe}^{2+} < \text{Al} < \text{Br}^-$ (c) $\text{Al} < \text{Br}^- < \text{Fe}^{2+}$ (d) $\text{Al} < \text{Fe}^{2+} < \text{Br}^-$
- Name the type of cell which was used in Apollo space programme.
(a) Primary cell (b) dry cell (c) fuel cell (d) mercury cell
- Actinoids exhibit greater number of oxidation states than lanthanoids. The main reason being
(a) more energy difference between $5f$ and $6d$ than between $4f$ and $5f$ orbitals.
(b) $4f$ orbitals are more diffused than the $5f$ orbitals
(c) lesser energy difference between $5f$ and $6d$ than between $4f$ and $5d$ orbitals
(d) more reactive nature of the actinoids than the lanthanoids.

For questions number 15 to 18, two statements are given - one labelled Assertion (A) and the other labelled Reason (R). Select the correct answer to these questions from the codes (i), (ii), (iii) and (iv) as given below:

- Both Assertion and Reason are correct statements; Reason (R) is correct explanation of Assertion (A).
- Both Assertion and Reason are correct statements, Reason (R) is not correct explanation of Assertion (A).

Chem Inquiry Practice With Transition Metals Answer Key

S Marginson



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Table of Contents Chem Inquiry Practice With Transition Metals Answer Key

1. Understanding the eBook Chem Inquiry Practice With Transition Metals Answer Key
 - The Rise of Digital Reading Chem Inquiry Practice With Transition Metals Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Chem Inquiry Practice With Transition Metals Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chem Inquiry Practice With Transition Metals Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chem Inquiry Practice With Transition Metals Answer Key
 - Personalized Recommendations
 - Chem Inquiry Practice With Transition Metals Answer Key User Reviews and Ratings
 - Chem Inquiry Practice With Transition Metals Answer Key and Bestseller Lists

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 - Highlighting and Note-Taking Chem Inquiry Practice With Transition Metals Answer Key
 - Interactive Elements Chem Inquiry Practice With Transition Metals Answer Key
8. Staying Engaged with Chem Inquiry Practice With Transition Metals Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chem Inquiry Practice With Transition Metals Answer Key
9. Balancing eBooks and Physical Books Chem Inquiry Practice With Transition Metals Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chem Inquiry Practice With Transition Metals Answer Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chem Inquiry Practice With Transition Metals Answer Key
 - Setting Reading Goals Chem Inquiry Practice With Transition Metals Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chem Inquiry Practice With Transition Metals Answer Key
 - Fact-Checking eBook Content of Chem Inquiry Practice With Transition Metals Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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