

Analysis of Vinegar

	Trial 1***	Trial 2	Trial 3
Volume of Vinegar used in titration.	<u>10 mL</u>	<u>10 mL</u>	<u>10 mL</u> (1 pt each)
Initial Buret Reading NaOH	<u>0 mL</u>	<u>0 mL</u>	<u>0 mL</u>
Final Buret Reading NaOH	<u>31.81 mL</u>	<u>31.87 mL</u>	<u>31.82 mL</u>
Total Volume NaOH added	_____	_____	_____
Acetic Acid Molarity	_____	_____	_____ (4 pts each)
Average Molarity of Acetic Acid in Vinegar	_____		

***Calculations Trial #1 (4 pts.)

Using the molarity of acetic acid in vinegar, calculate the mass percent of acetic acid in the original sample. Assume the density of vinegar is 1.0 g/mL . The formula for acetic acid is $\text{C}_2\text{H}_4\text{O}_2$. (5 pts.) (Hint: You start with molarity (units of $\frac{\text{mole}_{\text{C}_2\text{H}_4\text{O}_2}}{1 \text{ L}_{\text{vinegar}}}$) and you want to end with units of $\frac{\text{g}_{\text{C}_2\text{H}_4\text{O}_2}}{\text{g}_{\text{vinegar}}}$. You will multiply this result by 100 to get percent.) _____

Chem 11lab Manual Answers Experiment 12

Jo Allan Beran



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