

## Chemical Bonding

### SECTION 3

**SHORT ANSWER** Answer the following questions in the space provided.

- A** The notation for sodium chloride, NaCl, stands for one
  - formula unit.
  - molecule.
  - crystal.
  - atom.
- D** In a crystal of an ionic compound, each cation is surrounded by a number of
  - molecules.
  - positive ions.
  - dipoles.
  - negative ions.
- B** Compared with the neutral atoms involved in the formation of an ionic compound, the crystal lattice that results is
  - higher in potential energy.
  - lower in potential energy.
  - equal in potential energy.
  - unstable.
- B** The lattice energy of compound A is greater in magnitude than that of compound B. What can be concluded from this fact?
  - Compound A is not an ionic compound.
  - It will be more difficult to break the bonds in compound A than those in compound B.
  - Compound B has larger crystals than compound A.
  - Compound A has larger crystals than compound B.
- B** The forces of attraction between molecules in a molecular compound are generally
  - stronger than the attractive forces among formula units in ionic bonding.
  - weaker than the attractive forces among formula units in ionic bonding.
  - approximately equal to the attractive forces among formula units in ionic bonding.
  - equal to zero.
- Describe the force that holds two ions together in an ionic bond.

**Very strong overall attraction between positive and negative charges**

- What type of energy best represents the strength of an ionic bond?

**Lattice energy**

- What types of bonds are present in an ionic compound that contains a polyatomic ion?

# Chapter 6 Chemistry Test Answer Key

**Sir Thomas Edward Thorpe**



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Lab Experiments Modern Chemistry Harold Clark Metcalfe,1982  
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