

# Chapter 5 Bonding Test A

Do Not Write on This Test!!! You may use a periodic table.

**Multiple Choice**-Identify the choice that best completes the statement or answers the question.

- Electrons involved in bonding between atoms are
  - valence electrons
  - inside the nucleus
  - closest to the nucleus
  - positively charged
- Each family in the periodic table has its own characteristic properties based on the number of
  - neutrons
  - valence electrons
  - protons
  - ions
- What is the greatest number of valence electrons an atom can have? With the exception of helium,
  - 2
  - 3
  - 8
  - 12
- If atoms of a halogen nonmetal (Group 17) gains one electron, the atoms then have
  - no valence electrons
  - 7 valence electrons
  - 8 valence electrons
  - 17 valence electrons
- When an atom loses an electron, it becomes a
  - positive ion
  - negative ion
  - neutral ion
  - neutral atom
- An ionic bond is the attraction between
  - similarly charged ions
  - oppositely charged ions
  - neutral ions
  - neutral atoms
- The element boron is directly above aluminum on the periodic table. Which statement about boron is true?
  - Boron is in the same period as aluminum
  - Boron is in the same group as aluminum and has 5 valence electrons
  - Boron has 5 valence electrons and is in the same period as aluminum
  - Boron is in the same group as aluminum and has 3 valence electrons
- Magnesium bromide is an ionic compound with the chemical formula  $MgBr_2$ . What does the "2" tell you?
  - Bromide has a 2- charge.
  - There are two magnesium ions to every bromide ion.
  - There are two bromide ions for every magnesium ion.
  - Bromide has a 2+ charge.
- What is the chemical name for the compound with the formula  $Na_2S$ ?
  - sodium fluoride
  - magnesium sulfide
  - lithium oxide
  - sodium sulfide
- In the chemical formula for an ionic compound, which item is written first?
  - positive ion
  - negative ion
  - subscript
  - charge
- Which of the following is a characteristic property of ionic compounds?
  - They have low melting points.

# Chapter 5 Assessment For Chemistry

**Donna Larson**



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