

Use your notes from the Atomic Structure program to answer the following questions.

- The atomic number tells the number of positively charged protons in the nucleus of an atom. The atom is neutral because this is also the number of negatively charged electrons in the atom.
- The mass number tells the total number of protons and neutrons in the nucleus of an atom. These particles collectively are called nucleons since both are located in the nucleus.
- Isotopes are atoms of the same element with different numbers of neutrons which results in different mass numbers.
- Write the hyphen notation of the three isotopes of hydrogen:
 - Hydrogen - 1 ${}^1_1\text{H}$
 - Hydrogen - 2 ${}^2_1\text{H}$
 - Hydrogen - 3 ${}^3_1\text{H}$
- Fill in the following chart (Hint: Isotopes may have different masses than the periodic table lists!):

| Symbol | Atomic # | Mass # | # of protons | # of electrons | # of neutrons |
|--------|----------|--------|--------------|----------------|---------------|
| Hg | 80 | 201 | 80 | 80 | 121 |
| Mo | 42 | 96 | 42 | 42 | 54 |
| Cu | 29 | 64 | 29 | 29 | 35 |
| C | 6 | 14 | 6 | 6 | 8 |
| | 15 | 31 | 15 | 15 | 16 |
| Pb | 82 | 208 | 82 | 82 | 126 |
| Na | 11 | 24 | 11 | 11 | 13 |

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Bud Morgan



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