

## Chapter 6 Practice Test: Chemical Bonding

In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question.

- \_\_\_\_\_ 1. The charge on an ion is  
a. always positive.  
b. always negative.  
c. either positive or negative.  
d. zero.
- \_\_\_\_\_ 2. According to the octet rule, a calcium atom has a tendency to  
a. lose one electron.  
b. lose two electrons.  
c. gain one electron.  
d. gain two electrons.
- \_\_\_\_\_ 3. If a compound forms by ionic bonding, which is *not* true?  
a. A positively charged atom or group of atoms attracts a negatively charged atom or group of atoms.  
b. The net charge of the compound is zero.  
c. The compound contains just two atoms, each of opposite charge.  
d. Several ions group together in a tightly packed structure.
- \_\_\_\_\_ 4. The only property listed that is *not* characteristic of ionic compounds is  
a. high melting point.  
b. hardness.  
c. lack of crystal structure.  
d. good conductor of electricity.
- \_\_\_\_\_ 5. Which formula listed below represents a polyatomic ion?  
a.  $\text{NH}_4^+$   
b.  $\text{H}_2\text{SO}_4$   
c.  $\text{NaCl}$   
d.  $\text{H}_2\text{O}$
- \_\_\_\_\_ 6. The melting points of ionic compounds are higher than the melting points of molecular compounds because  
a. ionic substances tend to vaporize at room temperature.  
b. ionic substances are brittle.  
c. attractive forces between ions are greater than the attractive forces between molecules.  
d. the numbers of positive and negative charges are equal in an ionic compound.

# Chapter 6 Bonding Answer Key

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**Chemistry** John A. Olmsted, Gregory M. Williams, Robert Charles Burk, 2016-01-14 Olmsted Burk is an introductory general chemistry text designed specifically with Canadian professors and students in mind A reorganized Table of Contents and inclusion of SI units IUPAC standards and Canadian content designed to engage and motivate readers distinguish this text from many of the current text offerings It more accurately reflects the curriculum of most Canadian institutions Instructors will find the text sufficiently rigorous while it engages and retains student interest through its accessible language and clear problem solving program without an excess of material that makes most text appear daunting and redundant

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representation of the 6 5 species is obtained by a linear combination of three complete bond diagrams as in A1 which describe the optimal linear combination of 1 MOVB Configuration Wavefunctions CW s By the same token a total VB diagrammatic representation of the 6 6 species can be obtained by writing a dot structure as in B1 and taking this to mean the optimal linear combination of all VB CW s Next we can approximate the MOVB wavefunction of the 6 5 species by one complete or detailed bond diagram A2 No simple VB representation analogy can be given in this case Alternatively we can approximate the MOVB wavefunction by a linear combination of compact bond diagrams as in A3 in the way described before

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### Table of Contents Chapter 6 Bonding Answer Key

1. Understanding the eBook Chapter 6 Bonding Answer Key
  - The Rise of Digital Reading Chapter 6 Bonding Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 6 Bonding Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chapter 6 Bonding Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 6 Bonding Answer Key
  - Personalized Recommendations
  - Chapter 6 Bonding Answer Key User Reviews and Ratings
  - Chapter 6 Bonding Answer Key and Bestseller Lists

5. Accessing Chapter 6 Bonding Answer Key Free and Paid eBooks
  - Chapter 6 Bonding Answer Key Public Domain eBooks
  - Chapter 6 Bonding Answer Key eBook Subscription Services
  - Chapter 6 Bonding Answer Key Budget-Friendly Options
6. Navigating Chapter 6 Bonding Answer Key eBook Formats
  - ePub, PDF, MOBI, and More
  - Chapter 6 Bonding Answer Key Compatibility with Devices
  - Chapter 6 Bonding Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chapter 6 Bonding Answer Key
  - Highlighting and Note-Taking Chapter 6 Bonding Answer Key
  - Interactive Elements Chapter 6 Bonding Answer Key
8. Staying Engaged with Chapter 6 Bonding Answer Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chapter 6 Bonding Answer Key
9. Balancing eBooks and Physical Books Chapter 6 Bonding Answer Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chapter 6 Bonding Answer Key
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Chapter 6 Bonding Answer Key
  - Setting Reading Goals Chapter 6 Bonding Answer Key
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chapter 6 Bonding Answer Key
  - Fact-Checking eBook Content of Chapter 6 Bonding Answer Key
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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