



## CHAPTER 5 The Periodic Law

### SECTION 1 History of the Periodic Table

#### OBJECTIVES

1. Explain the roles of Mendeleev and Moseley in the development of the periodic table
2. Describe the modern periodic table.
3. Explain how the periodic law can be used to predict the physical and chemical properties of elements.
4. Describe how the elements belonging to a group of the periodic table are interrelated in terms of atomic elements.

### SECTION 2 Electron Configuration and the Periodic Table

#### OBJECTIVES

1. Describe the relationship between electrons in sublevels and the length of each period of the periodic table
2. Locate and name the four blocks of the periodic table. Explain the reason for these names.
3. Discuss the relationship between group configurations and group numbers.
4. Describe the location in the periodic table and the general properties of the alkali metals, alkaline-earth metals, the halogens, the chalogens, and the noble gases.

### SECTION 3 Electron Configuration and Periodic Properties

#### OBJECTIVES

1. Define atomic and ionic radii, ionization energy, electron affinity, and electronegativity.
2. Compare the periodic trends of atomic radii, ionization energy, and electronegativity, and state the reasons for these variations.
3. Define valence electrons, and state how many are present in atoms of each main-group element.
4. Compare the atomic radii, ionization energies, and electronegativities of the d-block elements with those of the main-group elements.

# Chapter 5 Test The Periodic Law

**David Baud**



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