

1. **Answer:** B

**Explanation:** We can see that both nuclides have different mass numbers and atomic numbers. However, we can determine that both nuclides have 16 neutrons. By definition, atoms with the same number of neutrons but different numbers of protons are called isotones.

2. **Answer:** A

**Explanation:** Remember that there are only two rules in balancing nuclear reactions, and these are the total of the subscripts and superscripts on the left side of the reaction must be equal to the total of the subscripts and superscripts on the right side. On the left side, the total of the superscripts is 29. To balance this, A must be 28. For the subscripts, the total on the left side is 14; therefore Z must be 13. With this, options C and D are already eliminated as the correct answer. Between options A and B, recall that the atomic number (Z) is unique for every element and can be seen in the periodic table as the number listed on top of the chemical symbol. With this, Al is the only element with an atomic number of 13, hence the correct answer is option A. Most examinations provide a copy of the periodic table of elements, but to be sure, it is advisable to memorize some of the common elements, their chemical symbols, atomic mass, and atomic number.

3. **Answer:** B

**Explanation:** By balancing the given reaction, we can determine that the correct identity of  ${}^A_ZX$  is actually  ${}^4_2\text{He}$ . Since an alpha particle was emitted, the reaction is an alpha decay or alpha particle emission reaction.

4. **Answer:** A

**Explanation:** If no values were given, you can always assign arbitrary values to visualize the problem. For this problem, let's say we have 1.0 g of  ${}^{131}\text{I}$  as the initial amount. After 8 days ( $1^{st}$  half-life), the amount will be reduced to half the original amount, which is 0.5 g. After another 8 days ( $2^{nd}$  half-life), the amount will be further reduced to half, from 0.5 g to 0.25 g. At this point,

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**Claire Aldridge, Karl Lee**



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