

CHAPTER 4

REACTIONS IN AQUEOUS SOLUTIONS

Problem Categories

Biological: 4.99, 4.102, 4.117, 4.145, 4.154, 4.169, 4.170, 4.171.

Conceptual: 4.7, 4.8, 4.11, 4.13, 4.17, 4.18, 4.58, 4.105, 4.121, 4.123, 4.129, 4.157, 4.158.

Descriptive: 4.9, 4.10, 4.12, 4.14, 4.21, 4.22, 4.23, 4.24, 4.33, 4.34, 4.43, 4.44, 4.51, 4.52, 4.53, 4.54, 4.55, 4.56, 4.103, 4.104, 4.117, 4.124, 4.125, 4.126, 4.129, 4.130, 4.131, 4.132, 4.133, 4.134, 4.135, 4.136, 4.137, 4.138, 4.139, 4.141, 4.142, 4.147, 4.149, 4.151, 4.152, 4.153, 4.155, 4.156, 4.159, 4.161, 4.163, 4.164, 4.165.

Environmental: 4.96, 4.131, 4.172.

Industrial: 4.147, 4.149, 4.150, 4.164.

Organic: 4.100, 4.112, 4.145, 4.154.

Difficulty Level

Easy: 4.7, 4.8, 4.9, 4.10, 4.11, 4.12, 4.17, 4.18, 4.19, 4.20, 4.48, 4.58, 4.63, 4.74, 4.75, 4.76, 4.142, 4.151, 4.153, 4.157.

Medium: 4.13, 4.14, 4.21, 4.22, 4.23, 4.24, 4.31, 4.32, 4.33, 4.34, 4.43, 4.44, 4.45, 4.46, 4.47, 4.49, 4.50, 4.51, 4.52, 4.53, 4.54, 4.55, 4.56, 4.57, 4.61, 4.62, 4.64, 4.65, 4.66, 4.67, 4.68, 4.70, 4.73, 4.84, 4.89, 4.90, 4.91, 4.92, 4.95, 4.96, 4.97, 4.98, 4.99, 4.100, 4.103, 4.104, 4.105, 4.107, 4.108, 4.109, 4.110, 4.111, 4.112, 4.113, 4.119, 4.120, 4.121, 4.122, 4.123, 4.124, 4.125, 4.126, 4.130, 4.131, 4.132, 4.133, 4.134, 4.135, 4.137, 4.139, 4.140, 4.141, 4.143, 4.144, 4.145, 4.147, 4.148, 4.149, 4.150, 4.152, 4.155, 4.156, 4.158, 4.160, 4.163, 4.167, 4.168.

Difficult: 4.69, 4.77, 4.78, 4.81, 4.82, 4.83, 4.101, 4.102, 4.106, 4.114, 4.115, 4.116, 4.117, 4.118, 4.127, 4.128, 4.129, 4.136, 4.138, 4.146, 4.154, 4.159, 4.161, 4.162, 4.164, 4.165, 4.166.

- 4.7 (a) is a strong electrolyte. The compound dissociates completely into ions in solution.
(b) is a nonelectrolyte. The compound dissolves in water, but the molecules remain intact.
(c) is a weak electrolyte. A small amount of the compound dissociates into ions in water.
- 4.8 When NaCl dissolves in water it dissociates into Na^+ and Cl^- ions. When the ions are hydrated, the water molecules will be oriented so that the negative end of the water dipole interacts with the positive sodium ion, and the positive end of the water dipole interacts with the negative chloride ion. The negative end of the water dipole is near the oxygen atom, and the positive end of the water dipole is near the hydrogen atoms. The diagram that best represents the hydration of NaCl when dissolved in water is choice (c).
- 4.9 Ionic compounds, strong acids, and strong bases (metal hydroxides) are strong electrolytes (completely broken up into ions of the compound). Weak acids and weak bases are weak electrolytes. Molecular substances other than acids or bases are nonelectrolytes.
- (a) very weak electrolyte (b) strong electrolyte (ionic compound)
(c) strong electrolyte (strong acid) (d) weak electrolyte (weak acid)
(e) nonelectrolyte (molecular compound - neither acid nor base)
- 4.10 Ionic compounds, strong acids, and strong bases (metal hydroxides) are strong electrolytes (completely broken up into ions of the compound). Weak acids and weak bases are weak electrolytes. Molecular substances other than acids or bases are nonelectrolytes.
- (a) strong electrolyte (ionic) (b) nonelectrolyte
(c) weak electrolyte (weak base) (d) strong electrolyte (strong base)
- 4.11 Since solutions must be electrically neutral, any flow of positive species (cations) must be balanced by the flow of negative species (anions). Therefore, the correct answer is (d).

Chapter 4 Electronic Communication System

Daniela Niemeyer



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