

ANSWERS

Nuclear Chemistry Review Questions

1. Compare the mass, charge, and penetration power of alpha, beta and gamma radiation.

	Mass	Charge	Penetration
Alpha	4	+	Lowest
Beta	0	-	Middle
Gamma	0	0	Highest

2. In 5.49 seconds, 1.20g of argon-35 decays to leave only 0.15g. What is the half-life?

1.20g \rightarrow 0.60g \rightarrow 0.30g \rightarrow 0.15g , therefore it took 3 half-life periods.

$$5.49 \div 3 = 1.83 \text{ sec}$$

3. How many days does it take 16.0g of gold-198 to decay to 1.0g?

16.0g \rightarrow 8.0g \rightarrow 4.0g \rightarrow 2.0g \rightarrow 1.0g , therefore 4 half-life periods

From Ref Table N, 1 half-life = 2.69d

$$2.69\text{d} \times 4 = 10.76 \text{ days}$$

4. Why does an atom undergo radioactive decay?

Because it has an unstable nucleus

5. Compare and contrast fission and fusion. Why is fission used for electrical and fusion is not?

Fission=Division, the splitting of large unstable atoms(nucleus)

Fusion=joining together, small isotopes of H, creates way more energy

6. How is nuclear chemistry applied to medicine? What type of radioisotope must be used when it is injected into a person?

Radioisotopes are used to treat and diagnose illnesses

If injected they must have short half-lives and quickly removed from the body

7. Write out the nuclear decay of radium-226.



8. What is different about nuclear reactions when compared to a chemical reaction?

Nuclear reactions transform atoms into other atoms, chemical reactions involved the rearrangement of electrons

9. At what point do all isotopes of an element become radioactive?

Up to and including atomic # 83

Chapter 2 study Guide Nuclear Chemistry Answers

John McMurry, Susan McMurry



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Study Guide with Answers to Selected Problems H. Stephen Stoker, Danny V. White, 2003-04 This supplement includes for each chapter a brief overview activities and practice problems to reinforce skills and a practice test The answers section includes answers for all odd numbered end of chapter exercises Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual Leo J. Malone, Theodore O. Dolter, 2012-01-03 The 9th edition of Malone s Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections assessment exercises at the end each section and relevant chapter problems at the end of each chapter A new Math Check allows quick access to the needed basic skill The first chapter now includes brief introductions to several fundamental chemical concepts and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem Every concept in the text is clearly illustrated with one or more step by step examples Making it Real essays have been updated to present timely and engaging real world applications emphasizing the relevance of the material they are learning This edition continues the end of chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter

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Study Guide, Answer Book to Accompany Organic Chemistry Alan S. Wingrove, Robert L. Caret, 1981

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