

# Ionic Exchange

The goal of this experiment is to understand the concept of the cation and anion exchange ion-exchange resin. It includes the following examples:  $\text{Na}^+$ ,  $\text{Ca}^{2+}$ ,  $\text{Mg}^{2+}$ . You should write down each of the changes to water.

Reverse ion-exchange resin is used to remove hardness. Hardness ions (calcium and magnesium) are replaced with sodium ions. The cation  $\text{Na}^+$  is called the "exchange ion" and  $\text{Ca}^{2+}$  is called the "target ion". Ion-exchange resin is used to remove hardness. The target ion is the hardness ion ( $\text{Ca}^{2+}$ ). The exchange ion is sodium ion ( $\text{Na}^+$ ).

## Critical Thinking Questions

1. Write the chemical formula for the cation and anion of the following salts:  $\text{NaCl}$ ,  $\text{CaCl}_2$ ,  $\text{MgCl}_2$ ,  $\text{Na}_2\text{SO}_4$ ,  $\text{CaSO}_4$ ,  $\text{MgSO}_4$ ,  $\text{Na}_2\text{CO}_3$ ,  $\text{CaCO}_3$ ,  $\text{MgCO}_3$ ,  $\text{Na}_2\text{HPO}_4$ ,  $\text{CaHPO}_4$ ,  $\text{MgHPO}_4$ ,  $\text{Na}_2\text{HPO}_4$ ,  $\text{CaHPO}_4$ ,  $\text{MgHPO}_4$ .

## Write equations for the following and show products.

1. Write the chemical formula for the cation and anion of the following salts:  $\text{NaCl}$ ,  $\text{CaCl}_2$ ,  $\text{MgCl}_2$ ,  $\text{Na}_2\text{SO}_4$ ,  $\text{CaSO}_4$ ,  $\text{MgSO}_4$ ,  $\text{Na}_2\text{CO}_3$ ,  $\text{CaCO}_3$ ,  $\text{MgCO}_3$ ,  $\text{Na}_2\text{HPO}_4$ ,  $\text{CaHPO}_4$ ,  $\text{MgHPO}_4$ ,  $\text{Na}_2\text{HPO}_4$ ,  $\text{CaHPO}_4$ ,  $\text{MgHPO}_4$ .

- $\text{Na}^+$  cation and  $\text{Cl}^-$  anion
- $\text{Ca}^{2+}$  cation and  $\text{Cl}^-$  anion
- $\text{Mg}^{2+}$  cation and  $\text{Cl}^-$  anion
- $\text{Na}^+$  cation and  $\text{SO}_4^{2-}$  anion
- $\text{Ca}^{2+}$  cation and  $\text{SO}_4^{2-}$  anion
- $\text{Mg}^{2+}$  cation and  $\text{SO}_4^{2-}$  anion
- $\text{Na}^+$  cation and  $\text{CO}_3^{2-}$  anion
- $\text{Ca}^{2+}$  cation and  $\text{CO}_3^{2-}$  anion
- $\text{Mg}^{2+}$  cation and  $\text{CO}_3^{2-}$  anion
- $\text{Na}^+$  cation and  $\text{HPO}_4^{2-}$  anion
- $\text{Ca}^{2+}$  cation and  $\text{HPO}_4^{2-}$  anion
- $\text{Mg}^{2+}$  cation and  $\text{HPO}_4^{2-}$  anion

The chemical formula of the cation and anion of the following salts:  $\text{NaCl}$ ,  $\text{CaCl}_2$ ,  $\text{MgCl}_2$ ,  $\text{Na}_2\text{SO}_4$ ,  $\text{CaSO}_4$ ,  $\text{MgSO}_4$ ,  $\text{Na}_2\text{CO}_3$ ,  $\text{CaCO}_3$ ,  $\text{MgCO}_3$ ,  $\text{Na}_2\text{HPO}_4$ ,  $\text{CaHPO}_4$ ,  $\text{MgHPO}_4$ ,  $\text{Na}_2\text{HPO}_4$ ,  $\text{CaHPO}_4$ ,  $\text{MgHPO}_4$ .

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**Yicheng Fang**



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