

Directions: Use the terms below to complete the following paragraphs about atoms, atomic mass, and isotopes. Terms may be used more than once.

six
 neutron(s)
 number
 proton(s)
 electrons
 mass
 isotopes
 quarks
 electron cloud
 six protons

The electron has very little mass compared to the 1. _____ or 2. _____. The mass of the atom depends on the nucleus and how many 3. _____ and 4. _____ it has. The sum of the protons and neutrons is the mass 5. _____ of an atom. The number of neutrons in an atom can be found by subtracting the atomic number from the 6. _____ number. The mass of the atom is so small that there is a measure called the atomic 7. _____ unit designated by amu. 8. _____ and 9. _____ make up the nucleus and are made up of 10. _____. There are 11. _____ uniquely different quarks. 12. _____ are found in an area around the nucleus called the 13. _____. The nuclei of all atoms of a given element always have the same number of 14. _____. They will also have the same number of 15. _____ around the nucleus. Some atoms may have more or fewer 16. _____ than will other atoms of the same element. Atoms of the same element with different numbers of neutrons are called 17. _____. Every atom of carbon must contain 18. _____ but some contain six neutrons and others have eight neutrons.

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