

**Directions:** Use the terms below to complete the following paragraphs about atoms, atomic mass, and isotopes. Terms may be used more than once.

six                      number                      electrons                      isotopes                      electron cloud  
 neutron(s)              proton(s)                      mass                      quarks                      six protons

The electron has very little mass compared to the 1. \_\_\_\_\_ or 2. \_\_\_\_\_. The mass of the atom depends on the nucleus and how many 3. \_\_\_\_\_ and 4. \_\_\_\_\_ it has. The sum of the protons and neutrons is the mass 5. \_\_\_\_\_ of an atom. The number of neutrons in an atom can be found by subtracting the atomic number from the 6. \_\_\_\_\_ number. The mass of the atom is so small that there is a measure called the atomic 7. \_\_\_\_\_ unit designated by amu. 8. \_\_\_\_\_ and 9. \_\_\_\_\_ make up the nucleus and are made up of 10. \_\_\_\_\_. There are 11. \_\_\_\_\_ uniquely different quarks. 12. \_\_\_\_\_ are found in an area around the nucleus called the 13. \_\_\_\_\_. The nuclei of all atoms of a given element always have the same number of 14. \_\_\_\_\_. They will also have the same number of 15. \_\_\_\_\_ around the nucleus. Some atoms may have more or fewer 16. \_\_\_\_\_ than will other atoms of the same element. Atoms of the same element with different numbers of neutrons are called 17. \_\_\_\_\_. Every atom of carbon must contain 18. \_\_\_\_\_ but some contain six neutrons and others have eight neutrons.

# Directed Reading For Content Mastery Section 1 The Nucleus

**Daniela Niemeyer**



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