

## Parts of the Atom

*\*Fill in the following table about the parts of the atom. (AMU stands for atomic mass unit.)*

PART OF THE ATOM	ELECTRIC CHARGE	LOCATION IN ATOM	ATOMIC MASS
proton	proton	Nucleus	2000 X More than an e <sup>-</sup>
neutron	neutron	Nucleus	2000 X More than an e <sup>-</sup>
electron	electron	Orbitals / Energy Levels	1

*\*Here's how to predict the structure of any element on the periodic table. For example, look at carbon:*

6
C
carbon
12.011

**Error: Reference source not found**atomic number: This is the number of protons (also the number of electrons)

**Error: Reference source not found**chemical symbol: This is an abbreviation for the element.

**Error: Reference source not found**element name: This is the name of the element.

**Error: Reference source not found**atomic mass: This number (rounded off) tells you the number of protons *plus* the number of neutrons. (Electrons are too small to be included.)

**To find the number of protons, look at the atomic number:**

Example for carbon: Atomic number: 6                      Number of protons: 6

**To find the number of neutrons, subtract the atomic number from the atomic mass:**

Example for carbon: Atomic mass: 12 (rounded)      Number of neutrons:  $12 - 6 = 6$

**To find the number of electrons, look at the atomic number:**

Example for carbon: Atomic number: 6                      Number of electrons: 6

*\*Complete the table for the elements with atomic numbers 1-10.*

ELEMENT NAME	ATOMIC NUMBER	ATOMIC MASS	PROTONS	NEUTRONS	ELECTRONS
Hydrogen	1	1.00	1	1	1
Helium	2	4.00	2	2	2
Lithium	3	6.91	3	3	3
Beryllium	4	9.01	4	4	4
Boron	5	10.81	5	5	5
Carbon	6	12.01	6	6	6
Nitrogen	7	14.00	7	7	7
Oxygen	8	15.99	8	8	8

# Directed Reading A The Parts Of An Atom Answers

**Rosina Ehmman**



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### **Table of Contents Directed Reading A The Parts Of An Atom Answers**

1. Understanding the eBook Directed Reading A The Parts Of An Atom Answers
  - The Rise of Digital Reading Directed Reading A The Parts Of An Atom Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Directed Reading A The Parts Of An Atom Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Directed Reading A The Parts Of An Atom Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Directed Reading A The Parts Of An Atom Answers
  - Personalized Recommendations
  - Directed Reading A The Parts Of An Atom Answers User Reviews and Ratings
  - Directed Reading A The Parts Of An Atom Answers and Bestseller Lists

5. Accessing Directed Reading A The Parts Of An Atom Answers Free and Paid eBooks
  - Directed Reading A The Parts Of An Atom Answers Public Domain eBooks
  - Directed Reading A The Parts Of An Atom Answers eBook Subscription Services
  - Directed Reading A The Parts Of An Atom Answers Budget-Friendly Options
6. Navigating Directed Reading A The Parts Of An Atom Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Directed Reading A The Parts Of An Atom Answers Compatibility with Devices
  - Directed Reading A The Parts Of An Atom Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Directed Reading A The Parts Of An Atom Answers
  - Highlighting and Note-Taking Directed Reading A The Parts Of An Atom Answers
  - Interactive Elements Directed Reading A The Parts Of An Atom Answers
8. Staying Engaged with Directed Reading A The Parts Of An Atom Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Directed Reading A The Parts Of An Atom Answers
9. Balancing eBooks and Physical Books Directed Reading A The Parts Of An Atom Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Directed Reading A The Parts Of An Atom Answers
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Directed Reading A The Parts Of An Atom Answers
  - Setting Reading Goals Directed Reading A The Parts Of An Atom Answers
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Directed Reading A The Parts Of An Atom Answers
  - Fact-Checking eBook Content of Directed Reading A The Parts Of An Atom Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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